

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the rates of the forward and reverse reactions are equal, and the concentrations of the reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of a system. For instance, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of either the reactants or the products. Equilibrium conditions are reached only at constant pressure, or concentration, or volume, or temperature. Le Chatelier's Principle states that applying stress changes the equilibrium of a system. A system at equilibrium will respond to stress in a way that minimizes the effect of the change and allows the system to reach a new equilibrium. When more products are added, an equilibrium reaction will shift to the left to consume the added products and reach a new equilibrium. When more reactants are added, the equilibrium will shift to the right to consume the added reactants and reach a new equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, pH, temperature, and catalyst changes in the following reactions. Equilibrium is reached when the rates of the forward and reverse reactions are equal, and the concentrations of the reactants and products remain constant. For example, if the concentration of a reactant is increased, the equilibrium will shift to the right to consume the added reactant and reach a new equilibrium.



Adding NaOH solution Fe³⁺ reacts with OH⁻ and this shifts Fe³⁺ from the reactant side, causing the forward concentration shift to reactant side to produce more reactant side a yellow color change.

Procedure (This Group):

- | | |
|--|---|
| 1. Prepare 100 mL of 0.1 M FeCl ₃ solution. | 1. Prepare 100 mL of 0.1 M KSCN solution. |
| 2. Prepare 100 mL of 0.1 M NaOH solution. | 2. Prepare 100 mL of 0.1 M HCl solution. |
| 3. Prepare 100 mL of 0.1 M FeCl ₃ solution. | 3. Prepare 100 mL of 0.1 M KSCN solution. |
| 4. Prepare 100 mL of 0.1 M NaOH solution. | 4. Prepare 100 mL of 0.1 M HCl solution. |
| 5. Prepare 100 mL of 0.1 M FeCl ₃ solution. | 5. Prepare 100 mL of 0.1 M KSCN solution. |
| 6. Prepare 100 mL of 0.1 M NaOH solution. | 6. Prepare 100 mL of 0.1 M HCl solution. |
| 7. Prepare 100 mL of 0.1 M FeCl ₃ solution. | 7. Prepare 100 mL of 0.1 M KSCN solution. |
| 8. Prepare 100 mL of 0.1 M NaOH solution. | 8. Prepare 100 mL of 0.1 M HCl solution. |

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