

CHEMISTRY FORM 4 SCHEMES OF WORK TERM 1,2,3

		SCHEME OF WORK		FORM FOUR CHEMISTRY		TERM ONE		
WK NO	LJ NO	TOPIC / SUBTOPIC	LESSON / SPECIFIC OBJECTIVES	TEACHING / LEARNING ACTIVITIES	MATERIALS / RESOURCES	REFERENCES	REMARKS	
1		ACIDS, BASES AND SALTS.	<u>By the end of the lesson, the learner should be able to:</u>					
	1.	Strength of acids.	Define an acid in terms of hydrogen ions.	Class experiments: investigate reactions of magnesium and zinc carbonate with different acids. Make and record observations in tabular form.	Magnesium strip, zinc carbonate, 2M HCl, 2M H ₂ SO ₄ , 2M ethanoic acid.	K.L.B. BK IV Pages 1-4		
	2	Acids in aqueous form.	Explain strength of acids in aqueous form in terms of number of hydrogen ions present.	Make deductions from the observations. Write relevant chemical equations and ionic equations. Detailed discussion leading to the definition of an acid and explanation of strength of an acid.				
	3.	pH values of acids.	Determine strength of acids using pH values.	Q/A: review determination of strength of acids using a litmus paper and pH scale. Class / group experiments: record colour of universal indicator in 2M HCl and 2M ethanoic acid.	Universal indicator, 2M HCl, 2M ethanoic acid, dry cells, carbon electrodes, milli-ammmeters, wires, switches etc.	K.L.B. BK IV Pages 4-6		
	4	Electrical conductivities of aqueous acids.	Determine strengths of acids by comparing their electrical conductivities. Classify acids as either strong or weak in terms of partial dissociations in aqueous solutions.	Set up voltmeters of 2M HCl and 2M ethanoic acid in turn. Record ammeter of current . Discuss the observations. Write corresponding ionic equations.				
	5	Definition of a base in terms of hydroxide ions.	Define a base in terms of hydroxide ions.	Teacher demonstration: Dissolve calcium hydroxide in water. Carry out litmus test on the resulting solution. Discuss the results; hence define a base in terms of hydroxide ions.	Red litmus paper, calcium hydroxide solid.	K.L.B. BK IV Pages 6-7		

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Scheme Of Work Of Chemistry For Third Term2015

Wolfgang Guggemos



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