

Average Atomic Mass

How can the masses of the particles within an atom be determined?

Objectives

After observing three gases, three ions, and naturally occurring three isotopes, the masses of these particles, ions, and atoms should be determined. Comparison of particles including those belonging to the same element should show three different values resulting from variation of particles used to construct these different particles. However, the relative mass of one of the particles will be used primarily to determine the other two. The relative mass of the particles will be determined by using the relative masses of the particles.

Model 1 - A Sample of Magnesium Metal



1. What are the masses of the three isotopes of magnesium shown in Model 1?
They are 24, 25, and 26.
2. What are the masses of the three isotopes of magnesium shown in Model 1?
They are 24, 25, and 26.
3. How are the masses of the three isotopes of magnesium shown in Model 1? They are 24, 25, and 26.
The masses of the three isotopes of magnesium are different, but they are all close to each other.
4. How are the masses of the three isotopes of magnesium shown in Model 1? They are 24, 25, and 26.
The masses of the three isotopes of magnesium are different, but they are all close to each other.

	^{24}Mg	^{25}Mg	^{26}Mg
Number of atoms	10	10	10

5. What are the masses of the three isotopes of magnesium shown in Model 1?
They are 24, 25, and 26.
 6. What are the masses of the three isotopes of magnesium shown in Model 1? They are 24, 25, and 26.
The masses of the three isotopes of magnesium are different, but they are all close to each other.
- As a group, determine the average mass of the three isotopes of magnesium shown in Model 1. The average mass of the three isotopes of magnesium is 24.3.

Periodicity Of Elements Pogil Key

Linda Trombley,J. Weston Walch (Firm)

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