

Periodic Trends

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. An element has the following properties:

- Atomic Number: 19
- Electronegativity: 0.9
- Ionization Energy: 418 kJ/mol

5. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

7. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

9. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

Name: _____ Date: _____ Period: _____

Periodic Trends Test Study Guide

Atomic Structure and Properties of the Periodic Table

1. What are the three subatomic particles found in the periodic table?

- Protons: positively charged, found in the nucleus.
- Neutrons: neutral, found in the nucleus.
- Electrons: negatively charged, found in the electron cloud.

2. Label the following parts of the periodic table: Main Groups, Transition Metals, Lanthanides, Actinides, Metalloids, Nonmetals, Noble Gases, Halogens, Alkali Metals, Alkaline Earth Metals, Boron Group, Carbon Group, Nitrogen Group, Oxygen Group, Halogen Group, Noble Gas Group.

3. Complete the following table:

Group	1	2	13	14	15	16	17	18
Period 1	H	He						
Period 2	Li	Be	B	C	N	O	F	Ne
Period 3	Na	Mg	Al	Si	P	S	Cl	Ar
Period 4	K	Ca	Sc	Ti	V	Cr	Mn	Fe
Period 5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru
Period 6	Cs	Ba	La	Hf	Ta	W	Re	Os
Period 7	Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs

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Period 5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru
Period 6	C							

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