

Science Fusion
Grade 5
Homeschool Pacing Guide

Unit 1: How Scientists Work

Options for Instruction: Two parallel paths meet the unit objectives, with a strong inquiry strand woven into each. Follow the Print Path, the Digital Path, or your customized combination of print, digital, and inquiry.

Note: Many of the Hands-on Inquiry activities can be completed using common, household objects. However, some do require more specialized scientific equipment. Please check the materials list in the TE.

Pacing Guide			
SE = Student Edition Interactive Worktext			TE = Teacher Edition
Days	Activity Type	Print Path	Digital Path
Unit 1 Opener			
Lesson 1: What Do Scientists Do?			
1-2 days	Big Idea	SE/*TE, pp. 1-2	
	Lesson	SE/*TE, pp. 3-13	Screens 1-9
	Hands-on Inquiry		Inquiry Flip Chart, p. 2 *TE, p. 1D, 3A
1 day	Review	SE/*TE, pp. 14-16	
	Assessment		¹ Lesson 1 Quiz
Lesson 2: What Skills Do Scientists Use?			
1-2 days	Lesson	SE/*TE, pp. 17-23	Screens 1-8
	Hands-on Inquiry		Inquiry Flip Chart, p. 3 *TE, p. 1E, 17A
1 day	Review	SE/*TE, pp. 24-26	
	Assessment		¹ Lesson 2 Quiz

Science Fusion Homeschool Pacing Guide

**Michael A. Dispezio, Marjorie
Frank, Michael R. Heithaus, Donna Ogle**

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web limiting reactant theoretical and percentage yields key $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 120.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{100.26 \text{ g}}{120.3 \text{ g}} \times 100 = 83.3\%$

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