



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use permanent marker for writing and drawing.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2 _____

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fus} for 1-hexanol in kJ mol⁻¹ and in $^{\circ}\text{C}$ respectively.

Compound	Molecular mass (M _r)	Boiling point (T _b)
A	102	162
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? at 0 $^{\circ}\text{C}$ and 101 kPa of external pressure? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its relative value of these molecules. (2)
- 3.1.4 What kind of the compound A or B is 1-hexanol? Explain based on its molecular formula. (2)
- 3.1.5 What compound B or A is not the higher vapor pressure? (2)
- 3.1.6 What compound B or A is not the higher vapor pressure? (2)

Compound B (M_r 102) is shown in the table. It has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 162 $^{\circ}\text{C}$.

- 3.1.6 Explain why the boiling point of compound B is relatively higher than that of compound A and B by referring to the T_b and molecular mass of the molecules. (2)

3.2 Substances C is added to water. It is a pure substance. C is not water. It is added to water and the solution is shown in the diagram below.



- 3.2.1 Is substance C ionic or covalent, and the solution formed? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the solution in the intermolecular forces involved. (2)

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CL Gary

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