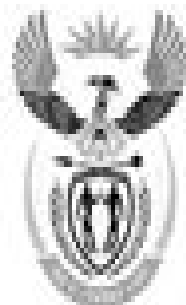


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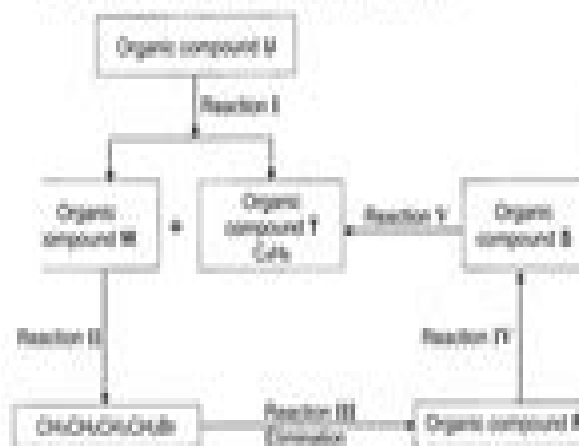
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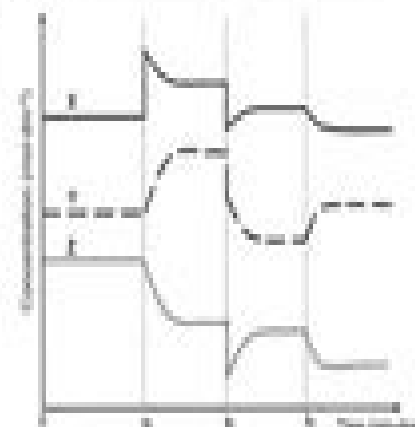
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Diagram on page 9 is redrawn below for easy reference.



Changes to the conditions of equilibrium are made at different times. The graph shows the results obtained. X, Y and Z represent the gases in the above reaction.



2.1.1 Define for both chemical equilibrium

Use the graph to answer the questions below.

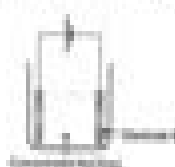
The inorganic reagents shown below are available for reactions IV and Y.

|                 |                                        |           |     |                |
|-----------------|----------------------------------------|-----------|-----|----------------|
| Br <sub>2</sub> | H <sub>2</sub> SO <sub>4</sub> (conc.) | NaOH(aq.) | HBr | H <sub>2</sub> |
|-----------------|----------------------------------------|-----------|-----|----------------|

4.5 Write down the TYPE of elimination in reaction III.

4.6 Compounds R and T are positional isomers.

4.7 Identify the reagent used in Reaction IV. (1)



4.8 Draw the half-cell reaction for the zinc electrode. (1)



4.9 Identify the reaction.

Write a balanced chemical equation for the reaction.

4.10 Draw a half-cell reaction for the zinc electrode. (1)

4.11 The simplified diagram below represents an electrochemical cell. The electrodes are made of copper.

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**J Ma**



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