

Chapter 6 Chemical Bonds

Section 6.2 Covalent Bonding**(pages 165–169)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy (page 165)

Relating Text and Visuals As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Molecular Models	
Model	Description
Electron dot	
Structural formula	
Space-filling	
Electron cloud	

Covalent Bonds (pages 165–167)

- Describe a covalent bond. _____

- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
 - electron dot
 - structural formula
 - space-filling
 - electron cloud
- Describe a molecule. _____

- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - zero
 - one
 - two
 - three

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K Morrison



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Table of Contents Physical Science Chapter 6 Chemical Bonds Chapter Review

1. Understanding the eBook Physical Science Chapter 6 Chemical Bonds Chapter Review
 - The Rise of Digital Reading Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Chemical Bonds Chapter Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Chemical Bonds Chapter Review

- Personalized Recommendations
- Physical Science Chapter 6 Chemical Bonds Chapter Review User Reviews and Ratings
- Physical Science Chapter 6 Chemical Bonds Chapter Review and Bestseller Lists
- 5. Accessing Physical Science Chapter 6 Chemical Bonds Chapter Review Free and Paid eBooks
 - Physical Science Chapter 6 Chemical Bonds Chapter Review Public Domain eBooks
 - Physical Science Chapter 6 Chemical Bonds Chapter Review eBook Subscription Services
 - Physical Science Chapter 6 Chemical Bonds Chapter Review Budget-Friendly Options
- 6. Navigating Physical Science Chapter 6 Chemical Bonds Chapter Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Chemical Bonds Chapter Review Compatibility with Devices
 - Physical Science Chapter 6 Chemical Bonds Chapter Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Highlighting and Note-Taking Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Interactive Elements Physical Science Chapter 6 Chemical Bonds Chapter Review
- 8. Staying Engaged with Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Chemical Bonds Chapter Review
- 9. Balancing eBooks and Physical Books Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Chemical Bonds Chapter Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Setting Reading Goals Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter 6 Chemical Bonds Chapter Review

- Fact-Checking eBook Content of Physical Science Chapter 6 Chemical Bonds Chapter Review
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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