

Periodic Trends Worksheet

Directions: Use the table to complete the following questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Na, Al, S, O

2. Write the following elements by increasing electronegativity: sulfur, oxygen, sodium, aluminum.

O, S, Al, Na

3. Which atom has the higher first ionization energy: Na or K?

Na has a higher first ionization energy than K because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been two cases each where a third element acts as a bridge between two groups.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium, cesium, barium, lanthanum, actinium, and the lanthanide series.

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 2, fluorine, neon, oxygen, nitrogen, carbon, boron, and beryllium.

8. Write the atoms in each group that have the largest atomic radius.

a. **Li** **Fr** **Bi**
 b. **Na** **K** **At**
 c. **Al** **Al** **At**
 d. **Si** **Si** **At**
 e. **P** **P** **At**
 f. **S** **S** **At**
 g. **Cl** **Cl** **At**
 h. **Br** **Br** **At**
 i. **I** **I** **At**
 j. **At** **At** **At**

k. **Fr** **Fr** **Fr**
 l. **Na** **Na** **Na**
 m. **Li** **Li** **Li**
 n. **Be** **Be** **Be**
 o. **B** **B** **B**
 p. **C** **C** **C**
 q. **N** **N** **N**
 r. **O** **O** **O**
 s. **F** **F** **F**
 t. **Ne** **Ne** **Ne**

9. Circle the atom in each group that has the greatest electronegativity.

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Katrin Zwirgmaier

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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