

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					18A	
	1A	2A	1A	2A	3A	4A	5A	6A	7A
1	H 1H	He 2He							
2	Li 3Li	Be 4Be							
3	Na 11Na	Mg 12Mg	Transition metals in between these rows				Al 13Al	Si 14Si	P 15P
4	K 19K	Ca 20Ca					Ga 31Ga	Ge 32Ge	As 33As

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Group 1A and Period 3?

What group and period is Helium in?

What element is in Group 2, and Period 10A? Element 2 is in Group 2A and Element 10A is in Group 10A.

What element is in Group 2A and Period 10?

What element Helium is?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Group 18, including Helium and Argon, has 8 valence electrons. How many valence electrons will they have?

How many electrons?

Hydrogen and Helium are both in Group 1A. Hydrogen 1 valence electron. Helium 2 valence electrons.

Group 18A: Helium has 2 valence electrons. Argon has 8 valence electrons. Krypton has 8 valence electrons. Xenon has 8 valence electrons. Radon has 8 valence electrons. Helium has 2 valence electrons. Argon has 8 valence electrons. Krypton has 8 valence electrons. Xenon has 8 valence electrons. Radon has 8 valence electrons.

Periods And Groups Valence Electrons Masses Answers

KJ Lindholm-Leary



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