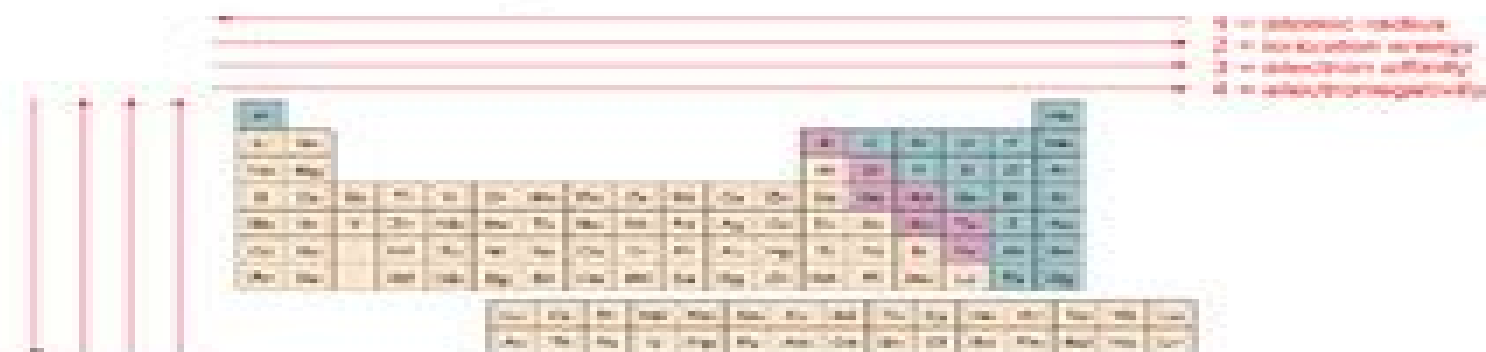


1. Match each definition with the correct trend below.

- |                            |  |
|----------------------------|--|
| <b>d</b> Ionization Energy | a. The energy change when one mole of electrons is added to one mole of gaseous atoms, forming negatively charged ions.                |
| <b>c</b> Atomic Radius     | b. A measure of an atom's ability to attract electrons in a chemical bond when forming a compound.                                     |
| <b>b</b> Electronegativity | c. The distance from the nucleus to the outermost electron in an atom, or half the distance between two nuclei in a diatomic molecule. |
| <b>a</b> Electron Affinity | d. The energy required to remove one mole of electrons from one mole of gaseous atoms.   |

2. Draw and label arrows on the periodic table below that indicate the direction of each trend.



3. Arrange the following elements in order of increasing atomic radius:

Be, Cl, F, I

F, Cl, Be, I

4. Rank the following elements in order of decreasing atomic radius:

C, Li, B, Be

Li, Be, B, C

5. Place the following elements in order of increasing ionization energy:

Li, K, F, O

K, Li, O, F

# Periodic Trends And Mixed Review

**Robert James Cimasi**



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