

# Periodic Trends Worksheet

Name: \_\_\_\_\_ Grade: \_\_\_\_\_ ID: \_\_\_\_\_

**Directions:** Use your notes to answer the following questions.

- 1 Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

**Ans:** O, C, Al, K

- 2 Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

**Ans:** Ne, Al, S, O

- 3 Why does fluorine have a higher ionization energy than iodine?

**Ans:** Fluorine has a smaller atomic radius, so the proton can exert a greater pull on electron (Valence shell is close)

- 4 Why do elements in the same family generally have similar properties?

**Ans:** They have the same valence shell electron arrangement.

- 5 Indicate whether the following properties increase or decrease from left to right across the periodic table.

- a. atomic radius (excluding noble gases) **Decrease**
- b. first ionization energy **Increase**
- c. electronegativity **Increase**

- 6 What trend in atomic radius occurs across the periodic table? What causes this trend?

**Ans:** Atomic radius increases down a group because energy levels (shells) are added.

- 7 What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

**Ans:** Ionization energy increases across a period because as elements become less metallic, it requires more energy to remove an electron.

- 8 Circle the atom in each pair that has the largest radius.

a. Al or B

d. O or F

b. Na or Al

e. Br or Cl

c. S or O

f. Mg or Ca

# Periodic Table Trends Answer Key

**Richard Bailey**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Periodic Table Trends Answer Key:**

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**Activities With Real-Life Applications** Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5 12 The book offers a dynamic collection of classroom ready lessons projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life

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**Chemistry in a Month** John Wilkinson, 2003

*Shriver and Atkins' Inorganic Chemistry* Peter Atkins, 2010 Inorganic Chemistry fifth edition represents an integral part of a student s chemistry education Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text culminating in the cutting edge research topics of the Frontiers which illustrate the dynamic nature of inorganic chemistry

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Neil D. Jespersen, Pamela Kerrigan, 2024-07-02

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## **Table of Contents Periodic Table Trends Answer Key**

1. Understanding the eBook Periodic Table Trends Answer Key
  - The Rise of Digital Reading Periodic Table Trends Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Trends Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Trends Answer Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Periodic Table Trends Answer Key
  - Personalized Recommendations
  - Periodic Table Trends Answer Key User Reviews and Ratings
  - Periodic Table Trends Answer Key and Bestseller Lists
5. Accessing Periodic Table Trends Answer Key Free and Paid eBooks
  - Periodic Table Trends Answer Key Public Domain eBooks
  - Periodic Table Trends Answer Key eBook Subscription Services
  - Periodic Table Trends Answer Key Budget-Friendly Options
6. Navigating Periodic Table Trends Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Table Trends Answer Key Compatibility with Devices
  - Periodic Table Trends Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Trends Answer Key
  - Highlighting and Note-Taking Periodic Table Trends Answer Key
  - Interactive Elements Periodic Table Trends Answer Key
8. Staying Engaged with Periodic Table Trends Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Trends Answer Key
9. Balancing eBooks and Physical Books Periodic Table Trends Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Trends Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Trends Answer Key
  - Setting Reading Goals Periodic Table Trends Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Periodic Table Trends Answer Key
  - Fact-Checking eBook Content of Periodic Table Trends Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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money loss during the period being reported an income statement represents a period of time as does the cash flow statement this contrasts with the balance sheet which represents a single moment in time

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