

CHAPTER 6

STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

~~octaves~~~~atomic mass~~~~atomic number~~~~size~~~~elements~~~~properties~~~~Henry Mosely~~~~eight~~

protons

periodic law

Dmitri Mendeleev

accepted

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Be	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Mosely. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

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Table of Contents Periodic Table Study Guide Answers Key

1. Understanding the eBook Periodic Table Study Guide Answers Key
 - The Rise of Digital Reading Periodic Table Study Guide Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Study Guide Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Study Guide Answers Key
 - Personalized Recommendations
 - Periodic Table Study Guide Answers Key User Reviews and Ratings
 - Periodic Table Study Guide Answers Key and Bestseller Lists
5. Accessing Periodic Table Study Guide Answers Key Free and Paid eBooks

- Periodic Table Study Guide Answers Key Public Domain eBooks
 - Periodic Table Study Guide Answers Key eBook Subscription Services
 - Periodic Table Study Guide Answers Key Budget-Friendly Options
6. Navigating Periodic Table Study Guide Answers Key eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Table Study Guide Answers Key Compatibility with Devices
 - Periodic Table Study Guide Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Table Study Guide Answers Key
 - Highlighting and Note-Taking Periodic Table Study Guide Answers Key
 - Interactive Elements Periodic Table Study Guide Answers Key
8. Staying Engaged with Periodic Table Study Guide Answers Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Study Guide Answers Key
9. Balancing eBooks and Physical Books Periodic Table Study Guide Answers Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Study Guide Answers Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Study Guide Answers Key
- Setting Reading Goals Periodic Table Study Guide Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Study Guide Answers Key
- Fact-Checking eBook Content of Periodic Table Study Guide Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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