

Percent Composition and Molecular Formula Worksheet Key

1. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? **C_3H_3O mass = 55 g/mole**
2. If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula? **$C_6H_6O_2$**
3. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? **Li_2CO_3**
4. If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula? **Li_2CO_3**

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. **$C_4O_2H_8$**
6. A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. **$C_8H_8O_2$**
7. A compound with an empirical formula of $CFBrO$ and a molar mass of 254.7 grams per mole. **$C_2F_2Br_2O_2$**
8. A compound with an empirical formula of C_2H_8N and a molar mass of 46 grams per mole. **C_2H_8N**

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. **CH_2O**
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? **$C_2H_4O_2$**
11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula? **C_6H_7N**
12. The molar mass of aniline is 93 g/mol. What is its molecular formula? **C_6H_7N**
13. Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C_2H_5NO . **%C 40.668 %H 8.533 %N 23.713 %O 27.086**
14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? **PCl_3**
15. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. **Hg_2O**
16. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % **C_7H_7NO**
17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %

$C_3H_7NO_3$ empirical formula

$C_3H_7NO_3$ molecular formula

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Nigel P. Freestone



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