

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					18A	
	1A	2A	13A	14A	15A	16A	17A	18A	18A
1	1H								2He
2	3Li	4Be							10Ne
3	11Na	12Mg							18Ar
4	19K	20Ca							36Kr

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Period 2, and Group 16A? Period 2 means 2 columns that is Column 16A. The element is oxygen.

What element is in Period 1 and Group 1?

What element is in Group 2A, and Period 3?

What group number is in Column 17?

What group number is in Column 16?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are both in Group 1A. Helium has 2 valence electrons. Hydrogen has 1 valence electron.

Group 1A, including H and Li, all have 1 valence electron. How many valence electrons will they have?

How many electrons?

Group 18A - These are noble gases that have a full shell of electrons. Elements are in Group 18 and they have 8 valence electrons. Helium has 2 valence electrons. "Ne" is a noble gas. Group 18A has a full shell of valence electrons. Helium has 2 valence electrons. Helium has 2 valence electrons. Helium has 2 valence electrons.

Periods And Groups Valence Electrons Masses Answers

SA Dillow



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