

Chapter 5 • Periodicity Activity**ATOMIC RADII****Directions:** Using the given data for the sizes of atoms (atomic radii) below,

1. Create a line graph (use a ruler) by graphing *Period 2's* data on the graph below.
2. In a *different color*, create a line graph *Period 3's* data on the same graph.
3. Label each axis, and include a key/legend.
4. Answer all questions.

	1	2	13	14	15	16	17	18
1	H 30							He -----
2	Li 123	Be 89	B 88	C 77	N 70	O 66	F 64	Ne -----
3	Na 157	Mg 136	Al 125	Si 117	P 110	S 104	Cl 99	Ar -----
4	K 203	Ca 174	Ga 125	Ge 122	As 121	Se 117	Br 114	Kr -----
5	Rb 216	Sr 192	In 159	Sn 149	Sb 141	Te 137	I 133	Xe -----
6	Cs 235	Ba 198	Tl 155	Pb 154	Bi 152	Po 153	At -----	Rn -----

Note: Mark the vertical axis from 0 – 160 (by 20's)

Trend in Atomic Size Across the Periodic Table

Li Na	Be Mg	B Al	C Si	N P	O S	F Cl

Horizontal Trend Questions

1. What is the trend in atomic size as you go *across* the periodic table? _____
2. This is true because there are more _____ (protons / electrons / layers of electrons) pulling the electron cloud in toward the nucleus.

Vertical Trend Questions

3. Examine the sizes of the Group 1 elements (H, Li, Na, K, Rb, Cs). What is the trend in atomic size as you go *down* a column of the periodic table? _____
4. This is true because there are more _____ (protons / electrons / layers of electrons) making the electron cloud larger.

Periodicity Activity Answers

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