

Name: \_\_\_\_\_

## Atomic Structure & the Periodic Table Study Guide

### Atomic Structure

| Particle | Charge             | Where are they found?                        |
|----------|--------------------|----------------------------------------------|
| Proton   | Positive           | Nucleus                                      |
| Neutron  | Neutral/ no charge | Nucleus                                      |
| Electron | negative           | Electron cloud/energy levels outside nucleus |

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**
2. What are the vertical columns on the periodic table called? Groups or families
  - a. How many of these are on the periodic table? **18**
  - b. Why are the elements in these columns? They have the same chemical properties (& same # of valence electrons)
  - c. Elements in the same group have the same number of valence electrons.
3. What are the horizontal rows on the periodic table called? periods
  - a. How many of these are on the periodic table? **7**
  - b. Elements in these rows have the same number of energy (*electron shells*) levels.
4. What are valence electrons? **Electrons in the outer energy level**
5. How many electrons are in each energy level?  
1<sup>st</sup> energy level **2**      2<sup>nd</sup> energy level **8**      3<sup>rd</sup> energy level **8**
6. The atomic number is equal to the # of protons and also # of electrons
7. The atomic mass is equal to the # protons + # of neutrons
8. How do you find the number of neutrons? Atomic mass – atomic number
9. We view the order of the periodic table today according to the atomic number

# Periodic Table Printable Study Guide

**JL Elias**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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## **Table of Contents Periodic Table Printable Study Guide**

1. Understanding the eBook Periodic Table Printable Study Guide
  - The Rise of Digital Reading Periodic Table Printable Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Printable Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Printable Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Printable Study Guide
  - Personalized Recommendations
  - Periodic Table Printable Study Guide User Reviews and Ratings

- Periodic Table Printable Study Guide and Bestseller Lists
- 5. Accessing Periodic Table Printable Study Guide Free and Paid eBooks
  - Periodic Table Printable Study Guide Public Domain eBooks
  - Periodic Table Printable Study Guide eBook Subscription Services
  - Periodic Table Printable Study Guide Budget-Friendly Options
- 6. Navigating Periodic Table Printable Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Table Printable Study Guide Compatibility with Devices
  - Periodic Table Printable Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Printable Study Guide
  - Highlighting and Note-Taking Periodic Table Printable Study Guide
  - Interactive Elements Periodic Table Printable Study Guide
- 8. Staying Engaged with Periodic Table Printable Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Printable Study Guide
- 9. Balancing eBooks and Physical Books Periodic Table Printable Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Table Printable Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Printable Study Guide
  - Setting Reading Goals Periodic Table Printable Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Printable Study Guide
  - Fact-Checking eBook Content of Periodic Table Printable Study Guide
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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