Gas Liquid Solid Unknown

Periodic Table

Reactions

Single Displacement → A + BC = B + AC Double Displacement → AB + CD = BC + AD

Synthesis -- A + B = AB

Decomposition → AB = A + B

Combustion → Hydrocarbon + O, = CO, + H,O Neutralization -- Acid + Base = Salt + H.O

He -

Ne

Kr.

-

Xe

-

Rn

-

Og

Li	Be
Na	Mg

н

Polyatomic ions (ion composed of multiple atoms)

Cr

Character

Mo

Mary Indiana, ma

Min

Manyorman

Tc

Technolium:

Re

Benedict

NO₋ → Nitrate (-1)

Ti

Station.

 Z_{S}

Dennison.

PO, → Phosphate (-3)

OH → Hydroxide (-1)

HCO_s → Bicarbonate (-1)

NO₂ → Nitrite (-1)

SO, → Sulfate(-2)

Ni

HILLS

Pd

Paladian

Pt

Co

Colonia

Rh

Blockett

le.

Million

CO, → Carbonate (-2)

NH, → Ammonium (+1)

Cu

Copper

Ag

Au.

Ro

ı	-	Sales	-	-	-	
	Al	Si	P	S	CI	A
ı	Atomina	1	200	1	-	-

As

Sb

Bi

Station	Hagemen	9115555
K	Са	Sc
Rb	Sr	Y
Cs	Ва	57-71
Fr	Ra	89-103

HI W Ta Technical Supplied Parl more

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Traffice.

Ge

Sn

Pb

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Br

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AL

Ts

Albani Metals

Decayly in lease bonds that to

Has a vigorous reaction when bonding with halogens, payour a

Alkaline Earth Metals Two saterops stectrons

Planetive but not as reactive as halogens and alkali metals because \$45.80 it needs to lose 2 electrons to have a full valence shell

Bonds with halocens, and other nero metals.

Transition Metals/Metals: Transition metals are metals that vary in valence electrons and can have multiple ion charges. They have similar qualities as alkali metals and alkaline earth metals: they can form metallic and ionic bonds. Metals with multiple ion charges are called multivalent metals. For example, copper has an ion charge of +1 and +2 so it is a multivalent metal. Therefore, the ion charge must be specified when written in word form for ionic bonds with multivalent metals. Metals can also bond with polystomic ions.

	La	Се	Pr	Nd	Pm	100000000000000000000000000000000000000	A 100 CO CO	Gd		1000000	0.000	Tm }		Lu
-	Ac	F }	Pa	U	Np		1000000	200000000000000000000000000000000000000	400	0.000	Fm	Md	No	Lr

mail and non-metal properties. They react with moters and no

y can not conduct electricity because they can not give off electrons, as cornert will not be able to go through ion mesule bond with metals in an ionic blind and other risk metals in a melecular bond. They can also bond with the

and metals. This is why notice gaser

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