

The Periodic Table

The Structure of the Periodic Table

- The periodic table displays the elements in order of increasing atomic number i.e. in order of increasing number of protons.
- Elements are arranged in 18 vertical columns called groups that are numbered 1 – 18.
- The seven horizontal rows in periodic table are called periods. The periods are numbered 1-7. The first period contains only two elements, hydrogen and helium.
- Elements within the same groups have similar chemical properties because they have the same number of valence electrons. The number of valence electrons is the same as the group number (in Roman numerals).

		GROUPS 1 - 18																					
PERIODS 1 - 7	1																	18					
	1	H 1.0079																	He 4.0026				
	2	Li 6.941	Be 9.0122															B 10.811	C 12.011	N 14.007	O 15.999	F 18.998	Ne 20.18
	3	Na 22.99	Mg 24.305															Al 26.982	Si 28.086	P 30.974	S 32.06	Cl 35.453	Ar 39.948
	4	K 39.098	Ca 40.078	Sc 44.956	Ti 47.88	V 50.942	Cr 51.996	Mn 54.938	Fe 55.847	Co 58.933	Ni 58.693	Cu 63.546	Zn 65.39	Ga 69.723	Ge 72.64	As 74.922	Se 78.96	Br 79.904	Kr 83.8				
	5	Rb 85.468	Sr 87.62	Y 88.906	Zr 91.224	Nb 92.906	Mo 95.94	Tc 97.907	Ru 101.07	Rh 102.91	Pd 106.42	Ag 107.87	Cd 112.41	In 114.82	Sn 118.71	Sb 121.76	Te 127.6	I 126.9	Xe 131.29				
	6	Cs 132.91	Ba 137.33	La 138.91	Hf 178.49	Ta 180.95	W 183.84	Re 186.21	Os 190.23	Ir 192.22	Pt 195.08	Au 196.97	Hg 200.59	Tl 204.38	Pb 207.2	Bi 208.98	Po 209	At 210	Rn 222				
7	Fr 223	Ra 226	Ac 227	Rf 261	Hs 277	Sg 294	Nh 294	Hl 294	Hs 294	Mt 294	Ds 294	Uub 294	Uut 294	Uuq 294	Uuh 294	Uus 294	Uuo 294						

Energy levels and period numbers

The electrons of an atom are arranged in energy levels. The period number tells us in which energy level an atom has its valence electrons.

For example:

Element	In which Group?	In which Period?	How many valence electrons	Which energy level are valence electrons?
Phosphorous	15	3	5	3 rd energy level
Potassium				
Calcium				
Oxygen				

Periodic Table Lessons Fourth Grade

L Darling-Hammond



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Table of Contents Periodic Table Lessons Fourth Grade

1. Understanding the eBook Periodic Table Lessons Fourth Grade
 - The Rise of Digital Reading Periodic Table Lessons Fourth Grade
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Lessons Fourth Grade
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Lessons Fourth Grade
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Lessons Fourth Grade
 - Personalized Recommendations
 - Periodic Table Lessons Fourth Grade User Reviews and Ratings
 - Periodic Table Lessons Fourth Grade and Bestseller Lists
5. Accessing Periodic Table Lessons Fourth Grade Free and Paid eBooks

- Periodic Table Lessons Fourth Grade Public Domain eBooks
 - Periodic Table Lessons Fourth Grade eBook Subscription Services
 - Periodic Table Lessons Fourth Grade Budget-Friendly Options
6. Navigating Periodic Table Lessons Fourth Grade eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Lessons Fourth Grade Compatibility with Devices
 - Periodic Table Lessons Fourth Grade Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Lessons Fourth Grade
 - Highlighting and Note-Taking Periodic Table Lessons Fourth Grade
 - Interactive Elements Periodic Table Lessons Fourth Grade
 8. Staying Engaged with Periodic Table Lessons Fourth Grade
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Lessons Fourth Grade
 9. Balancing eBooks and Physical Books Periodic Table Lessons Fourth Grade
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Lessons Fourth Grade
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Table Lessons Fourth Grade
 - Setting Reading Goals Periodic Table Lessons Fourth Grade
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Table Lessons Fourth Grade
 - Fact-Checking eBook Content of Periodic Table Lessons Fourth Grade
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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