

CHAPTER 6 STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, reads about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves

~~atomic mass~~

~~atomic number~~

~~size~~

~~elements~~

~~properties~~

~~Henry Moseley~~

~~eight~~

protons

~~periodic law~~

~~Dmitri Mendeleev~~

~~accepted~~

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Bo	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Moseley. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

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CL Gary



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Table of Contents Periodic Table And Periodic Law Solution Manual

1. Understanding the eBook Periodic Table And Periodic Law Solution Manual
 - The Rise of Digital Reading Periodic Table And Periodic Law Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table And Periodic Law Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Periodic Table And Periodic Law Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table And Periodic Law Solution Manual
 - Personalized Recommendations
 - Periodic Table And Periodic Law Solution Manual User Reviews and Ratings
 - Periodic Table And Periodic Law Solution Manual and Bestseller Lists

5. Accessing Periodic Table And Periodic Law Solution Manual Free and Paid eBooks
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 - Periodic Table And Periodic Law Solution Manual Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Periodic Table And Periodic Law Solution Manual
 - Highlighting and Note-Taking Periodic Table And Periodic Law Solution Manual
 - Interactive Elements Periodic Table And Periodic Law Solution Manual
8. Staying Engaged with Periodic Table And Periodic Law Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table And Periodic Law Solution Manual
9. Balancing eBooks and Physical Books Periodic Table And Periodic Law Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table And Periodic Law Solution Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table And Periodic Law Solution Manual
 - Setting Reading Goals Periodic Table And Periodic Law Solution Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table And Periodic Law Solution Manual
 - Fact-Checking eBook Content of Periodic Table And Periodic Law Solution Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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