

CECIL CONVENT SCHOOL

PERIODIC ASSESSMENT-2

Class 11 - Chemistry

Time Allowed: 1 hour

Maximum Marks: 20

General Instructions:

1. There are 11 questions in this question paper.
2. SECTION A - Q. No. 1 to 5 are very short answer questions carrying 1 mark each.
3. SECTION B - Q. No. 6 to 8 are short answer questions carrying 2 marks each.
4. SECTION C- Q. No. 9 to 11 is case based question carrying 3 marks each.
5. All questions are compulsory.
6. Attempt questions in order.

1. Water can be lifted into the water tank at the top of the house with the help of a pump. Then why is it not considered to be spontaneous? [1]
2. Neither q nor W is a state function but $q + W$ is a state function. Explain why? [1]
3. What is the enthalpy change for an adiabatic process? [1]
4. State the law of chemical equilibrium. [1]
5. A tank is full of water. Water is coming in as well as going out at same rate. What will happen to level of water in a tank? What is name given to such state? [1]
6. Calculate the standard enthalpy of formation of $\text{CH}_3\text{OH}(\text{l})$ from the following data: [2]



7. The equilibrium constant for a reaction is 10. What will be the value of $\Delta G^\ominus$? $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$, $T = 300 \text{ K}$. [2]
8. Write the equilibrium constant (K_c) expression for the following reactions. [2]



9. i. Why standard entropy of an elementary substance is not zero whereas standard enthalpy of formation is taken as zero? [3]
 ii. Under what conditions will the reaction occur, if
 a. both ΔH and ΔS are positive
 b. both ΔH and ΔS are negative
10. Dihydrogen gas is obtained from the natural gas by partial oxidation with steam as per following endothermic reaction: [1]
 $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$
 a. Write the expression for K_p for the above reaction.

Periodic Assessment 2 Chemistry Answers

Michael Brown



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Table of Contents Periodic Assessment 2 Chemistry Answers

1. Understanding the eBook Periodic Assessment 2 Chemistry Answers
 - The Rise of Digital Reading Periodic Assessment 2 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Assessment 2 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Periodic Assessment 2 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Assessment 2 Chemistry Answers
 - Personalized Recommendations
 - Periodic Assessment 2 Chemistry Answers User Reviews and Ratings
 - Periodic Assessment 2 Chemistry Answers and Bestseller Lists

5. Accessing Periodic Assessment 2 Chemistry Answers Free and Paid eBooks
 - Periodic Assessment 2 Chemistry Answers Public Domain eBooks
 - Periodic Assessment 2 Chemistry Answers eBook Subscription Services
 - Periodic Assessment 2 Chemistry Answers Budget-Friendly Options
6. Navigating Periodic Assessment 2 Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Assessment 2 Chemistry Answers Compatibility with Devices
 - Periodic Assessment 2 Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Assessment 2 Chemistry Answers
 - Highlighting and Note-Taking Periodic Assessment 2 Chemistry Answers
 - Interactive Elements Periodic Assessment 2 Chemistry Answers
8. Staying Engaged with Periodic Assessment 2 Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Assessment 2 Chemistry Answers
9. Balancing eBooks and Physical Books Periodic Assessment 2 Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Assessment 2 Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Assessment 2 Chemistry Answers
 - Setting Reading Goals Periodic Assessment 2 Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Assessment 2 Chemistry Answers
 - Fact-Checking eBook Content of Periodic Assessment 2 Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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