

## AP Questions: Thermochemistry

1984 B

| Substance                             | Standard Heat of Formation, $\Delta H_f^\circ$<br>in kJ/mol |
|---------------------------------------|-------------------------------------------------------------|
| C(s)                                  | 0.00                                                        |
| CO <sub>2</sub> (g)                   | -393.5                                                      |
| H <sub>2</sub> (g)                    | 0.00                                                        |
| H <sub>2</sub> O(l)                   | -285.83                                                     |
| O <sub>2</sub> (g)                    | 0.00                                                        |
| C <sub>3</sub> H <sub>7</sub> COOH(l) | ?                                                           |

The enthalpy change for the combustion of butyric acid at 25°C,  $\Delta H_{\text{comb}}$ , is -2,183.5 kilojoules per mole. The combustion reaction is



- From the above data, calculate the standard heat of formation,  $\Delta H_f^\circ$ , for butyric acid.
- Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

| Substance                           | Enthalpy of<br>Combustion, $\Delta H$<br>(kilojoules/mole) |
|-------------------------------------|------------------------------------------------------------|
| C(s)                                | -393.5                                                     |
| H <sub>2</sub> (g)                  | -285.8                                                     |
| C <sub>2</sub> H <sub>5</sub> OH(l) | -1366.7                                                    |
| H <sub>2</sub> O(l)                 | --                                                         |

- Write a separate, balanced chemical equation for the combustion of each of the following: C(s), H<sub>2</sub>(g), and C<sub>2</sub>H<sub>5</sub>OH(l). Consider the only products to be CO<sub>2</sub> and/or H<sub>2</sub>O(l).
- In principle, ethanol can be prepared by the following reaction:



Calculate the standard enthalpy change,  $\Delta H$ , for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard molar enthalpy of neutralization of a strong acid by a strong base. Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- What equipment would be needed?
- What measurements should be taken?
- Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, C<sub>3</sub>H<sub>8</sub>, is a hydrocarbon that is commonly used as fuel for cooking.

- Write a balanced equation for the complete combustion of propane gas, which yields CO<sub>2</sub>(g) and H<sub>2</sub>O(l).
- Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O<sub>2</sub> by volume.
- The heat of combustion of propane is -2,220.1 kJ/mol. Calculate the heat of formation,  $\Delta H_f^\circ$ , of propane given that  $\Delta H_f^\circ$  of H<sub>2</sub>O(l) = -285.3 kJ/mol and  $\Delta H_f^\circ$  of CO<sub>2</sub>(g) = -393.5 kJ/mol.
- Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 kilograms of water (specific heat = 4.18 J/g°C), calculate the increase in temperature of water.

# Objectives Questions On Thermochemistry

**H Kauffman**



## **Objectives Questions On Thermochemistry:**

The book delves into Objectives Questions On Thermochemistry. Objectives Questions On Thermochemistry is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Objectives Questions On Thermochemistry, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Objectives Questions On Thermochemistry
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- Chapter 3: Objectives Questions On Thermochemistry in Everyday Life
- Chapter 4: Objectives Questions On Thermochemistry in Specific Contexts
- Chapter 5: Conclusion

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## **Table of Contents Objectives Questions On Thermochemistry**

1. Understanding the eBook Objectives Questions On Thermochemistry
  - The Rise of Digital Reading Objectives Questions On Thermochemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Objectives Questions On Thermochemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Objectives Questions On Thermochemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Objectives Questions On Thermochemistry
  - Personalized Recommendations
  - Objectives Questions On Thermochemistry User Reviews and Ratings
  - Objectives Questions On Thermochemistry and Bestseller Lists
5. Accessing Objectives Questions On Thermochemistry Free and Paid eBooks
  - Objectives Questions On Thermochemistry Public Domain eBooks
  - Objectives Questions On Thermochemistry eBook Subscription Services
  - Objectives Questions On Thermochemistry Budget-Friendly Options
6. Navigating Objectives Questions On Thermochemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Objectives Questions On Thermochemistry Compatibility with Devices
  - Objectives Questions On Thermochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Objectives Questions On Thermochemistry
  - Highlighting and Note-Taking Objectives Questions On Thermochemistry
  - Interactive Elements Objectives Questions On Thermochemistry

8. Staying Engaged with Objectives Questions On Thermochemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Objectives Questions On Thermochemistry
9. Balancing eBooks and Physical Books Objectives Questions On Thermochemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Objectives Questions On Thermochemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Objectives Questions On Thermochemistry
  - Setting Reading Goals Objectives Questions On Thermochemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Objectives Questions On Thermochemistry
  - Fact-Checking eBook Content of Objectives Questions On Thermochemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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