

Full Name _____ Date _____ Per _____ Chemistry Chapter 6 The Periodic Table

Periodic Trends Practice

1. Identify each element as a metal, metalloid, or nonmetal

germanium _____ metalloid _____

zinc _____ metal _____

phosphorous _____ nonmetal _____

lithium _____ metal _____

2. Give two examples of elements for each category

noble gases _____ He, Ne, Ar, Kr, Xe, Rn _____

halogens _____ F, Cl, Br, I, At _____

alkali metals _____ Li, Na, K, Rb, Cs, Fr _____

alkaline earth metals _____ Be, Mg, Ca, Sr, Ba, Ra _____

transition metals _____ Any elements from 3B-28 _____

3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? Increases going down due to added energy levels. The inside energy levels are filled up so all remaining electrons

must go into open spaces further away from the nucleus.

4. What trend in atomic radius do you see as you go across a period/row on the periodic table? Decreases going across

due to electrons being added into the same valence shell. Protons would be added to the nucleus, making the positive charge stronger and electrons placed in the same space would be pulled in closer by the nucleus.

5. Circle the atom in each pair that has the largest atomic radius.

A. Al B

B. S O

C. Br Cl

D. Na Al

E. O F

F. Mg Ca

6. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

O < C < Al < K

7. Define ionization energy. The amount of energy required to remove an electron from a neutral atom

8. Is it easier to form a positive ion with an element that has a high ionization energy or an element that has a low ionization energy? Explain. It would be easier to form a + ion with a low ionization energy because less energy would need to be used to remove the electron.

9. What trend in ionization energy do you see as you go down a group/family on the periodic table? Going down a group the ionization energy decreases. The valence electrons are further away and easier to take away.

10. What trend in ionization energy do you see as you go across a period/row on the periodic table? What causes this trend? Going across the period, ionization energy increases. The atoms are smaller, so the nucleus is pulling electrons in closer making them harder to remove. The atoms are also increasing their number of valence electrons. As they get closer to having a complete set of 8 valence electrons they do not want to lose electrons any more, they want to gain them instead.

11. Circle the atom in each pair that has the greater ionization energy

A. Li Be

B. Ca Ba

C. Na K

D. P Ar

E. Cl Si

F. Li K

12. Define electronegativity. The ability of an atom to attract electrons in a bond.

13. What trend in electronegativity do you see as you go down a group/family on the periodic table? Values decrease within a group. Since the valence shell becomes further away, it is harder for the nucleus to draw electrons towards it.

14. What trend in electronegativity do you see as you go across a period/row on the periodic table? What causes this trend? Values increase across a period because the atoms are getting smaller and the nucleus can pull electrons in easier and the atoms are closer to being like noble gases. They want to be able to complete their valence shell.

Periodic Trends Summary Table Answer Key

JR Anderson



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Table of Contents Periodic Trends Summary Table Answer Key

1. Understanding the eBook Periodic Trends Summary Table Answer Key
 - The Rise of Digital Reading Periodic Trends Summary Table Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Summary Table Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Summary Table Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Summary Table Answer Key

- Personalized Recommendations
 - Periodic Trends Summary Table Answer Key User Reviews and Ratings
 - Periodic Trends Summary Table Answer Key and Bestseller Lists
5. Accessing Periodic Trends Summary Table Answer Key Free and Paid eBooks
- Periodic Trends Summary Table Answer Key Public Domain eBooks
 - Periodic Trends Summary Table Answer Key eBook Subscription Services
 - Periodic Trends Summary Table Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Summary Table Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Trends Summary Table Answer Key Compatibility with Devices
 - Periodic Trends Summary Table Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Trends Summary Table Answer Key
 - Highlighting and Note-Taking Periodic Trends Summary Table Answer Key
 - Interactive Elements Periodic Trends Summary Table Answer Key
8. Staying Engaged with Periodic Trends Summary Table Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Summary Table Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Summary Table Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Summary Table Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Summary Table Answer Key
- Setting Reading Goals Periodic Trends Summary Table Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Summary Table Answer Key

- Fact-Checking eBook Content of Periodic Trends Summary Table Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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