

POGIL: Periodic Trends in Atomic Properties

Information

Properties such as the size of an atom (atomic radius), the energy required to remove an electron from an atom (ionization energy), and the energy associated with the addition of an electron to a gaseous atom (electron affinity) can be understood in terms of the electron configuration of the atom and a competition between electron-nucleus attraction and electron-electron repulsion.

Preliminary Activity - Complete the orbital diagrams for the ground state configuration.

	1s	2s	2p		1s	2s	2p
H				C			
He				N			
Li				O			
Be				F			
B				Ne			

Model 1: Electron Ionization & Capture and Electrostatic Interactions

Reaction Equation for Electron Ionization: $X \rightarrow X^+ + e^-$ $\Delta E = 1^{\text{st}}$ ionization energy

Reaction Equation for Electron Capture: $X + e^- \rightarrow X^-$ $\Delta E = \text{electron affinity}$

Dominant inter-particle forces in helium:

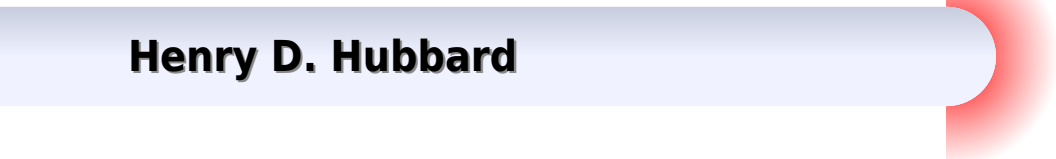


Dominant inter-particle forces in lithium:



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Henry D. Hubbard



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Science Spectrum Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004 **The Amazing World of the Atom and the Ultimate Periodic Table of the Elements** Costel Rizescu, Mihaela Rizescu, 2018-05-24 This book is addressed to undergraduate and graduate students taking physics and chemistry courses at an introductory level STEM high school students and to more experienced professors and researchers who need to stay up to date with the most recent advances in data as well. It is also beneficial to those involved in materials science, medical physics, biotechnology, astronomy, quantum chemistry, Earth and space science and other related fields. The volume consists of two parts: the first part is devoted to the quantum mechanics theory of atomic structures as discovered and verified by physics, whereas the second part presents and analyzes a large amount of the latest critically evaluated data for the properties of all elements on the periodic table of the elements. Periodic trends for ionization energy, electronegativity, atomic radius, density, and melting points are graphed and thoughtfully analyzed. Additional data of ionic radius, covalent radius, van der Waals radius, and 12 coordination atomic radius for almost all elements on the periodic table are also presented and evaluated. The authors discussed over 30 problems, review questions, and their solutions to help deepen the insight in this subject area to make this book a real study text. Plenty of problems are given to elucidate the material. This book represents the culmination of the authors' many years of research and teaching physics, materials science, and instrumentation. It is also a crystallization of their passion and strong interest in the history of atomic physics and the philosophy of science in general. From the beginning it was the authors' intention to include information in this book that is not easily located elsewhere or is not found completely in a single book or publication. Thus, while this book can be used as a text, the authors hope that it will be used as a useful reference too. Each chapter and subchapter of this book capitalizes on the strengths, comments, feedback, and criticism that the authors expect to have from students, faculty, and working professionals.

The periodic table. 4th ed, Atomic Properties of the Elements, 2001 Understanding the Atom: The Chemistry of the Noble Gases U.S. Atomic Energy Commission, 1967
Atomic Theories F. H. Loring, 1921 , *Chemistry and Our Universe*, Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure. Then witness what happens when three different alkali metals react with water. Theory forecasts a pronounced difference in the result. Is there Explanatory Key to the Periodic Chart of the Atoms Henry D. Hubbard, 1926 **The HISTORY of the ATOM the PERIODIC TABLE and RADIOACTIVITY** Alex Kimmons, 2021-05-15 THE HISTORY OF THE ATOM THE PERIODIC TABLE and RADIOACTIVITY monochrome Chemistry is an interesting and fundamental branch of science because it gives us the chance to explain the secrets of nature. What is water? What do we use in our cars as fuel? What is aspirin? What are perfumes made of? These kinds of questions and their answers are all part of the world of chemistry. There is no industry that does not depend upon chemical substances: the petroleum, pharmaceuticals, garment, aircraft, steel, and electronics industries, for example, as well as

agriculture all utilize the science of chemistry This book helps everyone to understand nature However one does not need to be a chemist or scientist to understand the simplicity within the complexity around us The aim was to write a modern up to date book where students and teachers can get concise information about the structure of substances Sometimes reactions are given in the detailed form but in general excessive detail has been omitted The book is designed to introduce fundamental knowledge in three areas the history of the atom the periodic table and radioactivity We will study the historical development of atomic structure theories the tendencies of elements in periods and groups and the types of emissions and uses of radioactivity

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