

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Table Trend Activities Answer Key

Didier Musso



Periodic Table Trend Activities Answer Key:

Hands-On General Science Activities With Real-Life Applications Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5-12. The book offers a dynamic collection of classroom ready lessons, projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life.

Chemistry in the Community. American Chemical Society, 2002. This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts.

Chemistry in the Community (Enhanced Core Four) American Chemical Society, 2006-02-15

Problems and Problem Solving in Chemistry Education Georgios Tsaparis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary, tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data, looking for patterns and making deductions. As an important higher order thinking skill, problem solving also constitutes a major research field in science education. Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving. The following situations are considered: some general others with a focus on specific areas of chemistry: quantitative problems, qualitative reasoning, metacognition and resource activation, deconstructing the problem solving process, an overview of the working memory hypothesis, reasoning with the electron pushing formalism, scaffolding organic synthesis skills, spectroscopy for structural characterization in organic chemistry, enzyme kinetics, problem solving in the academic chemistry laboratory, chemistry problem solving in context, team based active learning, technology for molecular representations, IR spectra simulation and computational quantum chemistry tools. The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry. With a foreword by George Bodner.

Science Units for Grades 9-12 Randy L. Bell, Joe Garofalo, 2005. Sample topics include cell division, virtual dissection, earthquake modeling, the Doppler Effect and more.

Chemistry in the Community (ChemCom) American Chemical Society, 2011-06-17. Touted as the most successful NSF funded project published, *Chemistry in the Community ChemCom* by the American Chemical Society (ACS) offers a meaningful and memorable chemistry program for all levels of high school students. ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios. Centered on decision making activities where students are responsible for generating data in an investigating, analyzing that data and then applying their chemistry knowledge to solve the presented problem. The text is intensively laboratory based with all 39 of the investigations integrated within the text, not separate from the reading. With the ChemCom program, students learn more organic and biochemistry, more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks, all within the relevance of solving problems that arise in

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applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress [Trends in Teaching Experimentation in the Life Sciences](#) Nancy J. Pelaez,Stephanie M. Gardner,Trevor R. Anderson,2022-05-11 This book is a guide for educators on how to develop and evaluate evidence based strategies for teaching biological experimentation to thereby improve existing and develop new curricula It unveils the flawed assumptions made at the classroom department and institutional level about what students are learning and what help they might need to develop competence in biological experimentation Specific case studies illustrate a comprehensive list of key scientific competencies that unpack what it means to be a competent experimental life scientist It includes explicit evidence based guidelines for educators regarding the teaching learning and assessment of biological research competencies The book also provides practical teacher guides and exemplars of assignments and assessments It contains a complete analysis of the variety of tools developed thus far to assess learning in this domain This book contributes to the growth of public understanding of biological issues including scientific literacy and the crucial importance of evidence based decision making around public policy It will be beneficial to life science instructors biology education researchers and science administrators who aim to improve teaching in life science departments Chapters 6 12 14 and 22 are available open access under a Creative Commons Attribution 4 0 International License via link [springer.com](https://www.springer.com) **Resources in Education** ,1995 [Fundamentals of Inorganic and Organic Chemistry](#) Dr. Priyanka Gupta Manglik,2024-08-10 Introduces the essential concepts of inorganic and organic chemistry including structure bonding reactions and their applications in science [Princeton Review AP Chemistry Premium Prep, 25th Edition](#) The Princeton Review,2023-08-01 EVERYTHING YOU NEED TO HELP SCORE A PERFECT 5 WITH THE BEST PRACTICE ON THE MARKET Equip yourself to ace the AP Chemistry Exam with this comprehensive study guide including 7 full length practice tests the MOST full length tests on the market thorough content reviews targeted strategies for every section and access to online extras Techniques That Actually Work Tried and true strategies to help you avoid traps and beat the test Tips for pacing yourself and guessing logically Essential tactics to help you work smarter not harder Everything You Need for a High Score Fully aligned with the latest College Board standards for AP Chemistry Comprehensive content review for all test topics Engaging activities to help you critically assess your progress Access to study plans a handy list of key equations helpful pre college information and more via your online Student Tools Premium Practice for AP Excellence 7 full length practice tests 5 in the book 2 online with detailed answer explanations Practice drills at the end of each content review chapter Review of important laboratory procedures and equipment **33 Years NEET Chapterwise & Topicwise Solved Papers CHEMISTRY (2020 - 1988) 15th Edition** Disha Experts, **Mathematics and Science Across the Curriculum**

,2002 **TEXT BOOK OF MEDICINAL CHEMISTRY-I** Dr. Prabhat Kumar Upadhyay, Dr. Narsukumari Korrapati, Dr Pichika Mallikarjuna Rao, Dr Balijepalli Madhu Katyayani, Dr. Mak Kit-Kay, Dr.Vishal Trivedi,2025-05-31 The Text Book of Medicinal Chemistry I is a foundational academic resource tailored for undergraduate pharmacy and life science students It begins with an insightful introduction to the history and development of medicinal chemistry followed by an exploration of essential physicochemical properties such as ionization solubility partition coefficient and bioisosterism that influence drug action The book delves into drug metabolism explaining both Phase I and Phase II reactions and emphasizing the stereochemical aspects affecting biotransformation A significant portion is dedicated to the autonomic nervous system detailing adrenergic and cholinergic neurotransmitters their biosynthesis metabolism receptors and drug interactions Comprehensive coverage is given to sympathomimetic agents and adrenergic antagonists including their classification mechanism of action and SAR The cholinergic section spans both parasympathomimetic agents and cholinergic blockers thoroughly explaining their pharmacological and structural aspects Sedatives and hypnotics such as benzodiazepines and barbiturates are analyzed for their therapeutic role and SAR Antipsychotic drugs including phenothiazine derivatives and newer analogs are presented with clarity The anticonvulsant chapter includes barbiturates hydantoins and newer agents like gabapentin and valproic acid along with mechanisms and SAR The text also explains the classification and use of general anesthetics including inhalation agents and dissociative anesthetics like ketamine In depth sections cover narcotic and non narcotic analgesics their SAR mechanism and therapeutic uses It emphasizes both opioid agonists and antagonists such as morphine and naloxone The final chapter provides detailed information on anti inflammatory drugs including NSAIDs like aspirin ibuprofen and diclofenac Throughout the book emphasizes structure activity relationships SAR mechanisms of action and pharmacological applications With drug examples and classifications well integrated this book provides a practical application oriented approach to medicinal chemistry It is designed to help students connect chemical structures with pharmacological effects Each chapter encourages critical thinking about how molecular features influence drug behavior The text is organized logically facilitating step by step learning of complex concepts Additionally it bridges the gap between basic science and clinical pharmacy making it ideal for academic and exam preparation **Princeton Review AP Chemistry**

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