

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

Name _____

Date _____

- 2 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 4 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 7.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 58.8 grams. Calculate the percent yield.

- 6 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). Calculate the theoretical yield.



What is the theoretical yield?

- 7 1.8 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 10.5 grams. Calculate the percent yield of phosphorus trichloride.

Name _____

Date _____

- 8 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 22.92 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 40.08 + 2(35.45) = 110.98 \end{array}$$

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 22.92 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \end{array}$$

- 9 The actual yield was 22.9 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{22.9 \text{ g}}{22.92 \text{ g}} \times 100 = 100.0\%$$

- 10 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 5.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 5.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 15.0 \text{ g KI} \end{array}$$

$$\begin{array}{l} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \end{array}$$

YIELD

ANSWER KEY INCLUDED

Percent Yield Example Answer Key

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Percent Yield Example Answer Key:

This book delves into Percent Yield Example Answer Key. Percent Yield Example Answer Key is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Percent Yield Example Answer Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Percent Yield Example Answer Key
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1 40 50 faustus speaks these lines near the end of his opening soliloquy in this speech he considers various fields of study one by one beginning with logic and proceeding through medicine and law

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4 of 5 over which country does faustus imagine himself as king prussia italy germany england 5 of 5 who lists several texts to help faustus study magic the evil angel

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