

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{54}_{26}\text{Fe}$	B2																
3 (b)	$^{23}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th><th>charge</th></tr></thead><tbody><tr><td>^{17}O</td><td>8</td><td>9</td><td>10</td><td>2-</td></tr><tr><td>^{27}Al</td><td>13</td><td>14</td><td>10</td><td>3+</td></tr></tbody></table>		protons	neutrons	electrons	charge	^{17}O	8	9	10	2-	^{27}Al	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
^{17}O	8	9	10	2-														
^{27}Al	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th></tr></thead><tbody><tr><td>^{32}S</td><td>16</td><td>16</td><td>16</td></tr><tr><td>$^{32}\text{S}^{2-}$</td><td>16</td><td>16</td><td>18</td></tr></tbody></table>		protons	neutrons	electrons	^{32}S	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
^{32}S	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2(\text{SO}_4)_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

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1. Understanding the eBook Ocr Examination Answers As Exam Cafe Chemistry
 - The Rise of Digital Reading Ocr Examination Answers As Exam Cafe Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Examination Answers As Exam Cafe Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Examination Answers As Exam Cafe Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Examination Answers As Exam Cafe Chemistry
 - Personalized Recommendations
 - Ocr Examination Answers As Exam Cafe Chemistry User Reviews and Ratings
 - Ocr Examination Answers As Exam Cafe Chemistry and Bestseller Lists
5. Accessing Ocr Examination Answers As Exam Cafe Chemistry Free and Paid eBooks
 - Ocr Examination Answers As Exam Cafe Chemistry Public Domain eBooks
 - Ocr Examination Answers As Exam Cafe Chemistry eBook Subscription Services
 - Ocr Examination Answers As Exam Cafe Chemistry Budget-Friendly Options
6. Navigating Ocr Examination Answers As Exam Cafe Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Examination Answers As Exam Cafe Chemistry Compatibility with Devices
 - Ocr Examination Answers As Exam Cafe Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Examination Answers As Exam Cafe Chemistry
 - Highlighting and Note-Taking Ocr Examination Answers As Exam Cafe Chemistry
 - Interactive Elements Ocr Examination Answers As Exam Cafe Chemistry
8. Staying Engaged with Ocr Examination Answers As Exam Cafe Chemistry

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Examination Answers As Exam Cafe Chemistry
9. Balancing eBooks and Physical Books Ocr Examination Answers As Exam Cafe Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Examination Answers As Exam Cafe Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Examination Answers As Exam Cafe Chemistry
- Setting Reading Goals Ocr Examination Answers As Exam Cafe Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Examination Answers As Exam Cafe Chemistry
- Fact-Checking eBook Content of Ocr Examination Answers As Exam Cafe Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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