

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reactants. These reactions are called reversible reactions, as the reaction can proceed in either direction. Once the equilibrium is reached the rates of the forward and reverse reactions are equal, and the concentrations of the reactants and products remain constant. The importance of understanding that equilibrium is a dynamic state is that it allows us to predict the behavior of a system. For example, if a reaction is at equilibrium, then the forward and reverse reactions continue to occur over time, but there is no net change in the concentrations of reactants and products. Equilibrium conditions are reached only at constant pressure, or concentration, or volume. Le Chatelier's Principle states that applying stress changes the equilibrium position of a system in order to minimize the effect of the change and bring the system back to equilibrium. When more products are added, an equilibrium reaction will shift to the left to consume the added products and bring the system back to equilibrium. When more reactants are added, the equilibrium will shift to the right to consume the added reactants and bring the system back to equilibrium. In this lab, you will be observing the changes in equilibrium when the amount of reactants, and temperature are changed. Follow Le Chatelier's Principle in the following reactions. Shifts in equilibrium are indicated by the color changes in solution according to the color indicators in the equations below. For example, if the color changes to red, the reaction shifts to the right to consume the added reactants and bring the system back to equilibrium.



Adding NaOH solution to $\text{Fe}^{3+}(\text{aq})$ and $\text{SCN}^{-}(\text{aq})$ will drive the reaction to the right, forming the complex ion FeSCN^{2+} . Adding the reactant Fe^{3+} will drive the reaction to the right, forming the complex ion FeSCN^{2+} .

Procedure (This Group):

- | | |
|---|--|
| 1. Prepare 100 mL of 0.1 M Fe^{3+} solution. | 2. Prepare 100 mL of 0.1 M SCN^{-} solution. |
| 3. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 4. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 5. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 6. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 7. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 8. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 9. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 10. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 11. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 12. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 13. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 14. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 15. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 16. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 17. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 18. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 19. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 20. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 21. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 22. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 23. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 24. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 25. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 26. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 27. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 28. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
| 29. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 30. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |
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| 99. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. | 100. Prepare 100 mL of 0.1 M FeSCN^{2+} solution. |

Observing Chemical Equilibrium Lab 46

Wolfgang Guggemos

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