CHEMISTRY Objective Final Exam Review 10-11

There are 50 questions on the test. The chapters are listed below. Each of the letters represents a concept/ idea that you should know or look up.

1. CH 6: Periodic trends, where do they increase & decrease
    a. Atomic radius & Ionic radius
    b. Electronegativity
    c. Electron affinity
    d. Ionization energy
    e. Reactivity of metals and of non-metals
2. CH 7: Look up properties of elements per group and the names of each group.
    a. Group 1 very reactive in water; group 7 most reactive non-metals etc.
    b. Fluorine most reactive in periodic table
    c. Group 1 will lose e- to become +1
    d. Elements in same group have same chem properties.
3. CH 8: Ionic compounds and bonds
    a. Why do atoms bond? How many electrons make them stable?
    b. How to name ionic compounds, go from name to formula
    c. What elements make up ionic compounds?
    d. Properties of ionic compounds
    e. What is a crystal lattice and the arrangement.
    f. What is metallic bonding?
4. CH 9: Covalent compounds
    a. What elements make up covalent compounds?
    b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
    c. What are pi bonds and sigma bonds?
    d. Bond length Vs. Bond strength, and Bond dissociation energy.
    e. What are molecules?
    f. What are the common diatomic molecules?
    g. Polar molecules, partial charges, what are they?
5. CH 10: Identify the types of chemical reactions
    a. Be able to write the products for a double replacement and single replacement reactions.
    b. Balancing equations
    c. Writing equations from word names.
6. CH 11: The mole
    a. Determine the molar mass of a compound
    b. Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
    c. Converting mole→ mole, mole → grams, mole → particles and visa versa
    d. Converting grams →mole, grams → grams, grams→ particles visa vera
7. CH 12:
    a. Stoichiometry conversions using balanced chemical equation
    b. Mole → mole, mole → grams, grams → grams

Objective Final Review For Chemistry

Rachel S Tattersall

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