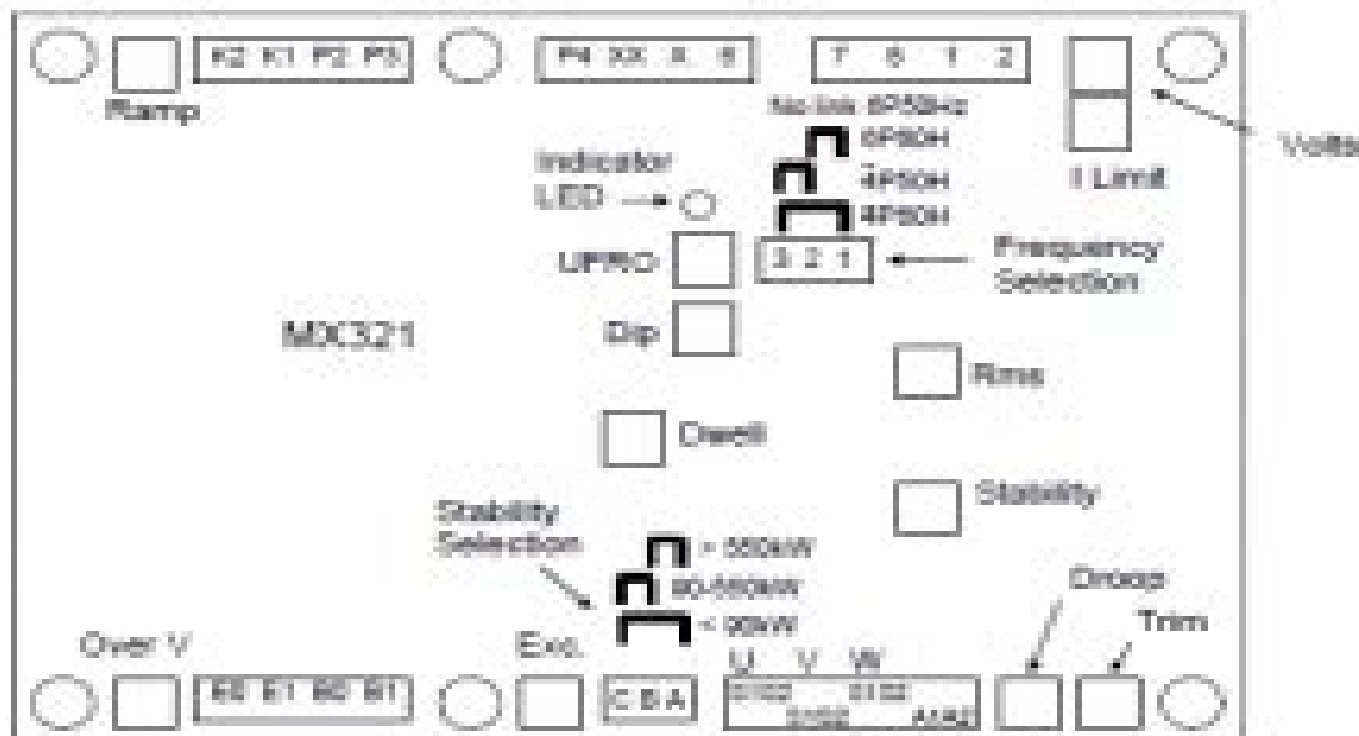


# FITTING AND OPERATING (Refer to generator wiring diagram for connection details)



## SUMMARY OF AVR CONTROLS

CONTROL	FUNCTION	DIRECTION
VOLTS	TO ADJUST GENERATOR OUTPUT VOLTAGE	CLOCKWISE INCREASES OUTPUT VOLTAGE
STABILITY	TO PREVENT VOLTAGE HUNTING	CLOCKWISE INCREASE THE DAMPING EFFECT
UPRO	TO SET THE UPRO KNEE POINT	CLOCKWISE REDUCES THE KNEE POINT FREQUENCY
DROOP	TO SET THE GENERATOR DROOP TO 3% AT OFF	CLOCKWISE INCREASES THE DROOP
TRIM	TO OPTIMISE ANALOGUE INPUT SENSITIVITY	CLOCKWISE INCREASES THE GAIN OR SENSITIVITY
EXC	TO SET THE OVER EXCITATION TRIP LEVEL	CLOCKWISE INCREASES THE TRIP LEVEL
DIP	TO SET THE Hz RELATED VOLTAGE DIP	CLOCKWISE INCREASES THE DIP
DWELL	TO SET THE Hz RELATED RECOVERY TIME	CLOCKWISE INCREASES THE RECOVERY TIME
I LIMIT	TO SET THE STATOR CURRENT LIMIT	CLOCKWISE INCREASES THE CURRENT LIMIT
OVER V	TO SET THE OVER VOLTAGE TRIP LEVEL	CLOCKWISE INCREASES THE TRIP LEVEL
RAMP	TO SET THE NO LOAD VOLTAGE RAMP UP TIME	CLOCKWISE INCREASES THE VOLTAGE RAMP TIME

# Newage Mx321 Avr Manual

**Peipei Pang**



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base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1  
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$   
 one mole of HCl would be fully neutralized by one mole of NaOH

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4$  aq  $2\text{NaOH}$  aq  $\text{Na}_2\text{C}_2\text{O}_4$   $4\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH}$  aq  $\text{C}_6\text{H}_4\text{COOH}$  aq  $\text{C}_6\text{H}_4\text{COONa}$  aq  $\text{H}_2\text{O}$  l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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