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 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2M HCl with the NaOH of unknown concentration in order to experimentally ascertain the concentration of the NaOH

this is a chemistry lab report on an acid base titration experiment - Dec 28 2021

web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction $NaOH$ aq C_6H_4COOH aq C_6H_4COONa aq H_2O l the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the