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web the balanced chemical equation for this reaction is  $\text{Mg}(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + \text{H}_2(\text{g})$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

[\$\text{HNO}\_3 + \text{Mg} \rightarrow \text{H}\_2\text{O} + \text{Mg}\(\text{NO}\_3\)\_2\$  chemical equation](#) - Aug 11 2023

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mg  $\text{hno}_3$   $\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{h}_2$  chemical equation balancer - Oct 13 2023

web magnesium nitric acid magnesium nitrate dihydrogen  $\text{mg}$   $\text{hno}_3$   $\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{h}_2$  is a single displacement substitution reaction where one mole of solid magnesium  $\text{mg}$  and two moles of aqueous nitric acid  $\text{hno}_3$  react to form one mole of aqueous magnesium nitrate  $\text{mg}$   $\text{no}_3$   $\text{2}$  and one mole of dihydrogen  $\text{h}_2$  gas show chemical

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web mar 4 2018  $\text{mg}$   $2\text{hno}_3$   $\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{h}_2$  on the left side there are  $\text{mg}$  1 magnesium in elemental form  $\text{h}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens  $\text{n}_2$  there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens  $\text{o}_6$

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web chemistry questions and answers 1 enter balanced molecular equation for the reaction between nitric acid and magnesium hydroxide express your answer as a chemical equation including phases 2 enter a net ionic equation for the reaction between nitric acid and magnesium hydroxide express your answer as a chemical equation including

$\text{mgco}_3$   $\text{hno}_3$   $\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{h}_2\text{o}$   $\text{co}_2$  chemical equation - Jul 10 2023

web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{mgco}_3$  and two moles of nitric acid  $\text{hno}_3$  react to form one mole of magnesium nitrate  $\text{mg}$   $\text{no}_3$   $\text{2}$  one mole of water  $\text{h}_2\text{o}$  and one mole of carbon dioxide  $\text{co}_2$  show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying  $\text{hno}_3$  by 2 we get  $\text{mg}$   $\text{s}$  magnesium  $2$   $\text{hno}_3$   $\text{l}$  nitric acid  $\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{s}$  magnesium nitrate  $\text{h}_2$   $\text{g}$  hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

$\text{mg}$   $\text{no}_3$   $\text{2}$   $\text{h}_2\text{so}_4$   $\text{mgso}_4$   $\text{hno}_3$  chemical equation - Apr 07 2023

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