

**FINAL**



**KWAZULU-NATAL PROVINCE**  
**EDUCATION**  
REPUBLIC OF SOUTH AFRICA

**NATIONAL  
SENIOR CERTIFICATE**

**GRADE 12**

**PHYSICAL SCIENCES P2 (CHEMISTRY)**

**COMMON TEST**

**JUNE 2024**

**MARKING GUIDELINES**

**MARKS: 100**

**These marking guidelines consist of 10 pages.**

# Physical Science Grade 12 Memo Paper June Exam2014

**RS Peters**



**Physical Science Grade 12 Memo Paper June Exam2014:**

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## **Physical Science Grade 12 Memo Paper June Exam2014** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Physical Science Grade 12 Memo Paper June Exam2014**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is  $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-\text{aq}$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver  $\text{I}^+$  ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13  $2\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$  because two  $\text{NH}_4^+\text{aq}$

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