

### Unit 3 Review - Atomic Structure & Periodic Table

#### Part 1: Answer the following questions.

1. a. Which scientist created the first modern atomic theory? John Dalton  
 b. What was his theory?
  1. All matter is made of extremely small particles called atoms.
  2. All atoms of a given element are identical and atoms of a specific element are different from any other element
  3. Atoms cannot be created, divided into smaller particles or destroyed
  4. Different atoms combine in simple whole number ratios to form compounds
  5. In a chemical reaction, atoms are separated, combined or rearranged.
- c. What was his model of the atom called? Billiard Ball
- d. What was proven wrong and why?
  1. All atoms of a given element are identical – wrong b/c of the existence of isotopes
  2. Atoms cannot be divided into smaller particles – wrong b/c of the existence of electrons, protons and neutrons
2. a. Which scientist did the gold foil experiment? Rutherford  
 b. What did he conclude from the gold foil experiment?
  1. Atoms are made up of mostly empty space
  2. Atoms have a small positive center call the nucleus
3. a. Who discovered the electron? J. J. Thomson  
 b. What was the name of his experiment? Cathode Ray Tube Experiment  
 c. What was his model of the atom called? Plum Pudding Model
4. What did Mendeleev do?  
Arranged the periodic table in order of increasing atomic mass
5. a. Who arranged the **modern** periodic table? Moseley  
 b. How is the modern periodic table arranged? Arranged in order of Atomic Number
6. a. What are the three subatomic particles? Give the symbols, charge, mass and location of each.

| Particles | Symbol | Charge   | Mass       | Location       |
|-----------|--------|----------|------------|----------------|
| Proton    | $p^+$  | Positive | 1 amu      | Nucleus        |
| Neutron   | $n^0$  | Neutral  | 1 amu      | Nucleus        |
| Electron  | $e^-$  | Negative | 1/2000 amu | Electron Cloud |

7. What is special about the proton? The proton IDENTIFIES THE ELEMENT
8. Fill in the blanks in the following:
  - a. Mass # =  $p^+ + n^0$
  - b. Atomic # = # of protons
9. What is the difference between Average Atomic mass and mass #? (Average) Atomic Mass is found on the periodic table and is the weighted average of the isotopes. The Mass Number is the mass of one element and must be given or calculated by adding the protons and neutron of that element.
10. Which of the following are isotopes of the same element? How do you know? A, and D. They have the same Atomic Number (# of Protons)
  - a.  $^{12}_5X$
  - b.  $^{15}_5X$
  - c.  $^{12}_6X$
  - d.  $^{15}_7X$
11. Calculate the average atomic mass for Boron if it has two naturally occurring isotopes: boron -10 (abundance of 19.8%) and boron - 11 (abundance of 80.2%).  

$$\{ 10 \times \frac{19.8}{100} \} + \{ 11 \times \frac{80.2}{100} \} = 10.802 \text{ amu}$$
12. Identify an element "X" by calculating the average atomic mass. X = 6 (abundance of 7.5%) and X= 7 (abundance 92.5%).  

$$\{ 6 \times \frac{7.5}{100} \} + \{ 7 \times \frac{92.5}{100} \} = 6.925 \text{ amu} \quad \text{Which is Lithium}$$

# Physical Science Atomic Structure Test Answers

**M Walker**



## **Physical Science Atomic Structure Test Answers:**

## Unveiling the Power of Verbal Beauty: An Psychological Sojourn through **Physical Science Atomic Structure Test Answers**

In some sort of inundated with screens and the cacophony of fast connection, the profound power and emotional resonance of verbal art frequently disappear in to obscurity, eclipsed by the constant assault of sound and distractions. Yet, nestled within the lyrical pages of **Physical Science Atomic Structure Test Answers**, a charming work of literary brilliance that impulses with organic emotions, lies an unforgettable journey waiting to be embarked upon. Penned by a virtuoso wordsmith, this interesting opus instructions viewers on a psychological odyssey, lightly revealing the latent potential and profound influence embedded within the complicated internet of language. Within the heart-wrenching expanse of the evocative analysis, we will embark upon an introspective exploration of the book is central styles, dissect its fascinating writing type, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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