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R Sandford



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web 7 Na + 2 C_2O_4 is the limiting reactant percent yield 86.6% 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + \text{C} \rightarrow \text{CO}_2 + \text{I}_2$ 80.0 grams of iodine(V) oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 \times 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

7.3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3