

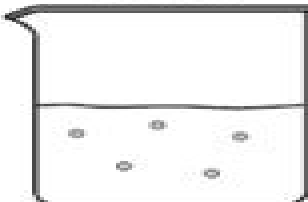
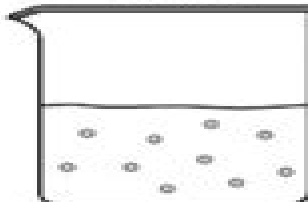
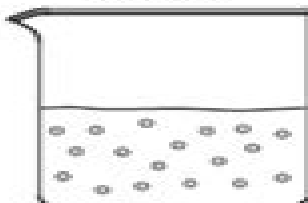
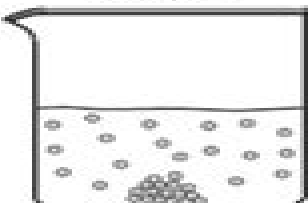
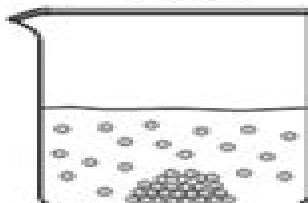
Saturated and Unsaturated Solutions

Is there a limit to the amount of solute that will dissolve in a solvent?

Why?

We use solutions every day. People who wear contact lenses use "lens solution" to rinse their contacts and keep them wet. Athletes who consume sports drinks after exercising benefit from the electrolytes in those solutions. This activity will explore whether or not there is a limit to how much of one substance can dissolve in another.

Model 1 – Saturated and Unsaturated Solutions

| <ul style="list-style-type: none"> All beakers contain 10.0 g of water. All beakers are kept at 20 °C. All solutions are stirred for 2 hours. Solute is the same substance in all beakers. | Unsaturated Solutions | |
|--|--|---|
| | <p>Beaker A</p>  <p>1.0 g of solute added</p> | <p>Beaker B</p>  <p>2.0 g of solute added</p> |
| | Number of dissolved particles | 5 |
| | Number of solid particles | 0 |
| Saturated Solutions | | |
| <p>Beaker C</p>  <p>3.6 g of solute added</p> | <p>Beaker D</p>  <p>7.0 g of solute added</p> | <p>Beaker E</p>  <p>9.0 g of solute added</p> |
| | | 18 |
| | 17 | |

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J Rink



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