(2)

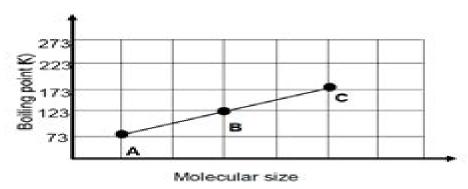
(1)

(3) [13]

#### QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The **letters A**, **B** and **C** represent the compounds CH<sub>4</sub>, C<sub>4</sub>H<sub>6</sub> and C<sub>3</sub>H<sub>8</sub> respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
  (2)
- 2.2 Consider the two molecules in the table below.

| NAME<br>OFSUBSTANCE | FORMULA         | MOLACULAR<br>MASS (g) | BOILING POINT |
|---------------------|-----------------|-----------------------|---------------|
| Ammonia             | NH <sub>3</sub> | 17                    | - 33          |
| Phosphine           | PH <sub>3</sub> | ×                     | - 87,4        |

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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