



OPERATORS MANUAL

FOR ROTAX® ENGINE TYPE 912 SERIES

Ref. No.: OM-912



ROTAX® 912 ULS 3 WITH OPTIONS

Rotax 912 Uls Manual

Y Pai



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web control of the i 2 c bus may be taken when the p bit is set or the bus is idle in firmware controlled master mode user code conducts all i 2 c bus operations based on start and stop bit condition detection start and stop condition detection is

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web lbidirectional i2c stands for inter integrated circuit communications i2c is implemented in the picmicro by a hardware module called the master synchronous serial port known as the mssp module

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example $\text{hcl} + \text{naoh} \rightarrow \text{nacl} + \text{h}_2\text{o}$

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$ eg 2 $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(aq)}$ you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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