

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equilibrium is at 1 bar and $T_f = 1.00^\circ C$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180.16 \text{ g}}}{0.100 \text{ kg}} = 5.55 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 5.55}{1000} = 0.0103^\circ C$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equilibrium is at 1 bar and $T_f = 1.00^\circ C$. $K_f = 1.86^\circ C$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60.06 \text{ g}}}{0.100 \text{ kg}} = 16.65 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 16.65}{1000} = 0.0310^\circ C$$

3. A 100 g mass of water at $100^\circ C$ is added to 100 g of water at $20^\circ C$. Assuming that the composition of the solution is homogeneous, $100^\circ C$ is the boiling-point of the solution (1 bar). What is the boiling-point of the solution (1 bar) if the composition is homogeneous? $T_f = 1.00^\circ C$.

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} \Rightarrow 0.0103^\circ C = \frac{1.86^\circ C \times \text{molality}}{1000}$$

$$\text{molality} = \frac{0.0103^\circ C \times 1000}{1.86^\circ C} = 5.55 \text{ mol/kg}$$

$$\Rightarrow \text{mol} = 5.55 \times 0.100 \text{ kg} = 0.555 \text{ mol}$$

$$m = 0.555 \text{ mol}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \Rightarrow 5.55 = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$\Rightarrow 0.555 \text{ mol} = \text{moles solute} \Rightarrow 0.555 \text{ mol}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.555 \text{ mol}}{0.100 \text{ kg}} = 5.55 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 5.55}{1000} = 0.0103^\circ C$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 1.00 g of water is composed. The freezing-point of this solution is $1.00^\circ C$. The normal freezing-point of pure water is $1.00^\circ C$ and the freezing-point depression constant for water is $1.86^\circ C$. Determine the molar mass of the compound. Assume ΔT_f is small. $T_f = 1.00^\circ C$.

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} \Rightarrow 0.0103^\circ C = \frac{1.86^\circ C \times \text{molality}}{1000} \Rightarrow \text{molality} = 5.55 \text{ mol/kg}$$

$$\text{molality} = \frac{\text{mol}}{\text{kg solvent}} \Rightarrow 5.55 = \frac{\text{mol}}{0.001 \text{ kg}}$$

$$\Rightarrow 0.00555 \text{ mol} = \text{mol of solute}$$

$$m = 0.00555 \text{ mol}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \Rightarrow 5.55 = \frac{\text{moles solute}}{0.001 \text{ kg}}$$

$$\Rightarrow 0.00555 \text{ mol} = \text{moles solute} \Rightarrow 0.00555 \text{ mol}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.00555 \text{ mol}}{0.001 \text{ kg}} = 5.55 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} \Rightarrow \frac{0.0103^\circ C \times 1000}{1.86^\circ C} = \text{molality}$$

$$\text{molality} = \frac{\text{moles solute}}{\text{kg solvent}} \Rightarrow 5.55 = \frac{\text{moles solute}}{0.001 \text{ kg}}$$

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