

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{54}_{26}\text{Fe}$	B2																
3 (b)	$^{24}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table><tr><td></td><td>protons</td><td>neutrons</td><td>electrons</td><td>charge</td></tr><tr><td><math>^{17}\text{O}</math></td><td>8</td><td>9</td><td>10</td><td>2-</td></tr><tr><td><math>^{27}\text{Al}</math></td><td>13</td><td>14</td><td>10</td><td>3+</td></tr></table>		protons	neutrons	electrons	charge	$^{17}\text{O}$	8	9	10	2-	$^{27}\text{Al}$	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
$^{17}\text{O}$	8	9	10	2-														
$^{27}\text{Al}$	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table><tr><td></td><td>protons</td><td>neutrons</td><td>electrons</td></tr><tr><td><math>^{32}\text{S}</math></td><td>16</td><td>16</td><td>16</td></tr><tr><td><math>^{32}\text{S}^{2-}</math></td><td>16</td><td>16</td><td>18</td></tr></table>		protons	neutrons	electrons	$^{32}\text{S}$	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
$^{32}\text{S}$	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2(\text{SO}_4)_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

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**KJ Lindholm-Leary**



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