

7. Chromite is the main ore of chromium. The chromium-containing compound in chromite is FeCr_2O_4 . The percentage of chromium in a sample of chromite can be determined using the method below.

Step 1

A 5.28 g sample of chromite ore is heated with sodium peroxide, Na_2O_2 .



Water is added to the resulting mixture.

Na_2O_2 dissolves in the water forming a solution containing CO_3^{2-} ions.

Step 2

The mixture from Step 1 is filtered and the filtrate is made up to 1.00 dm³ in a volumetric flask.

A 25.0 cm³ sample of this alkaline solution is pipetted into a conical flask and an excess of aqueous potassium iodide is added.

- A redox reaction takes place between I^- ions, CrO_4^{2-} ions and H_2O .
- In this reaction 1 mol CrO_4^{2-} reacts 1.5 mol I_2 .

Step 3

The resulting mixture is titrated with 0.100 mol dm⁻³ sodium thiosulfate, $\text{Na}_2\text{S}_2\text{O}_3$, to estimate the I_2 present.



The average titre of $\text{Na}_2\text{S}_2\text{O}_3$ was 26.5 cm³.

- (a) In Step 1, Na_2O_2 and NaFeO_2 react with water forming an alkaline solution containing a brown precipitate. This is not a redox reaction.

Write equations for:

- the reaction of Na_2O_2 with water
- the reaction of NaFeO_2 with water.



- (b) Calculate the percentage, by mass, of chromium in the ore.

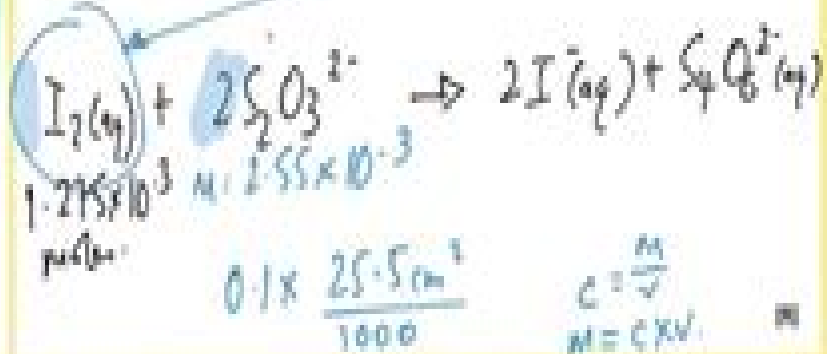
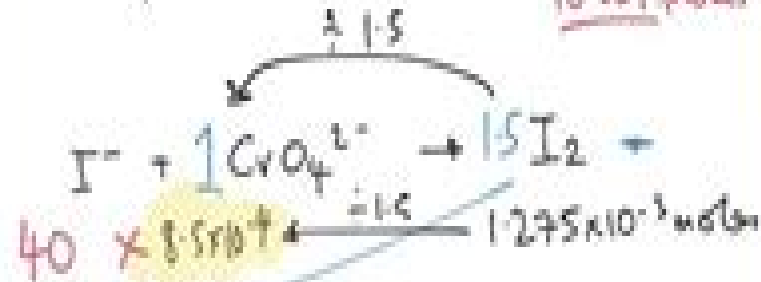
Give your answer to one decimal place.

$$\% \text{ Cr} = \frac{\text{mass of Cr}}{\text{mass of ore}} \times 100$$

$$\% \text{ Cr} = \frac{1.76}{5.28} \times 100 = 33.3\%$$



10.034 moles



- (c) The part refers to Step 2 of the method.

is the redox reaction between I^- ions, CrO_4^{2-} ions and H_2O .

- CrO_4^{2-} ions are reduced to chromate ions.
- I^- ions are oxidised to form I_2 .
- Construct an overall equation for the redox reaction and write half equations for the oxidation and reduction.

Overall equation:

Half equations:

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Padhraic Smyth



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