

- (b) A solution is prepared by the addition of 10.0 mL of 0.100 M CH_3COOH and 10.0 mL of 0.100 M NaOH .

- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

The total volume is 20.0 mL. The concentration of CH_3COOH is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a buffer: CH_3COOH and CH_3COO^- .

(d) The CH_3COOH and CH_3COO^- are 0.050 M and 0.050 M.

$$[\text{CH}_3\text{COOH}] = 0.050 \text{ M} + 0.050 \text{ M} = 0.100 \text{ M}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
Initial: 0.100 M, 0.100 M, 0.000 M, 0.000 M
Change: -0.050 M, -0.050 M, +0.050 M, +0.050 M
Equilibrium: 0.050 M, 0.050 M, 0.050 M, 0.050 M
The solution is a buffer solution of CH_3COOH and CH_3COO^- .
The concentration of CH_3COOH is 0.050 M.
The concentration of CH_3COO^- is 0.050 M.
The pH is 4.74.

- (b) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M and the concentration of CH_3COO^- is 0.050 M.

The pH is 4.74.



- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M and the concentration of CH_3COO^- is 0.050 M.

The pH is 4.74.

- (d) The pH is 4.74.

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

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$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

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$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \text{ M} \times 0.100 \text{ M}}{0.100 \text{ M}} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

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Rachel Sandford



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