

Boyle's Law

1. The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
2. Boyle's
3. Boyle's Law

Boyle's Theory and Boyle's Law

1. Boyle's of gases is caused by the elasticity of the gas molecules themselves that collide.
2. If the gas is compressible to small the volume is small which increases the pressure which is small.
3. Force exerts pressure on closed on the walls of the container.
4. Pressure of the gas will increase.

- Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. If the gas is released, it maintains a volume of 900 ml. above the pressure inside of 800 ml. What is the volume of the balloon at this altitude?

$$P_1 V_1 = P_2 V_2 \quad 800 \text{ ml.} \times 1 \text{ atm} = 900 \text{ ml.} \times P_2$$

$$P_2 = \frac{800 \text{ ml.} \times 1 \text{ atm}}{900 \text{ ml.}}$$

- Ex. 2. A gas at pressure of 100 mm. Hg. has a volume of 500 ml. If the pressure is increased until the volume is 400 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ mm. Hg.} \times 500 \text{ ml.} = P_2 \times 400 \text{ ml.}$$

$$P_2 = \frac{100 \text{ mm. Hg.} \times 500 \text{ ml.}}{400 \text{ ml.}}$$

ANSWER: 125 mm. Hg.

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