

Chemical Formula

Talking about proportions, molecules and relative of atoms indicated by a subscript.

Molecular $\text{C}_{10}\text{H}_{18}\text{O}_2$ (molecular)

Formula represents the number of _____ of _____
in a single molecule.

Empirical $\text{C}_5\text{H}_9\text{O}$ (molecular)

Formula represents the proportions of _____ of
positive and negative ions in one formula
unit.

Atomic Mass

The proportions that atoms combine through
ions together.

Masses

$\text{C}_{10}\text{H}_{18}\text{O}_2$ molecular

$\text{C}_5\text{H}_9\text{O}$ empirical charged
ions

Subscripts

$\text{C}_{10}\text{H}_{18}\text{O}_2$ molecular

$\text{C}_5\text{H}_9\text{O}$ empirical charged

$\text{C}_{10}\text{H}_{18}\text{O}_2$

Note Taking Guide Episode 601 Chemical Formula Key

J Dewey



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