

Worksheet #1: Significant Figures & Dimensional Analysis

Directions: For each problem below, write the equation and show your work. Be sure to **box your final answer**.

Part 1: Solve the following problems using scientific notation and rounding to the appropriate value.

- a. $300,000,000 \rightarrow 3.0 \times 10^8$
- b. $456,000 \rightarrow 4.6 \times 10^5$
- c. $0.000000000 \rightarrow 9.6 \times 10^{-9}$
- d. $-0.0000000 \rightarrow -3.6 \times 10^{-6}$

Part 2: Solve the following problems, and show your final answer with the appropriate number of significant figures.

- a. $1.20004 + 71.80 \rightarrow 73.00004 \approx 73.00$
- b. $98 + 9934.0 - 20 \rightarrow 10032 \approx 10030$
- c. $0.0007 + 2.15 + 0.0001 \rightarrow 2.1508 \approx 2.15$
- d. $4.5 \times 10^3 + 9.0 \times 10^2 + 89.22 \times 10^1 \rightarrow 4500 + 900 + 892.2 \rightarrow 6292.2 \approx 6300$
- e. $12.00 + 3.0 \times 10^3 - 0.7 \rightarrow 3012.3 \approx 3012$

Part 3: For each problem below, show your work. Always use units and box in your final answer.

- a. The density of pure silver is 10.5 g/cm^3 at 20°C . If 5.25 g of pure silver pellets are added to a graduated cylinder containing 11.2 mL of water, to what volume level will the water in the cylinder rise?

$d = \frac{m}{V}$

$d = 10.5 \text{ g/cm}^3 \quad 1 \text{ mL} = 1 \text{ cm}^3$

$m = 5.25 \text{ g}$

$V = ?$

$11.2 + 0.5 = 11.7 \text{ mL}$

$10.5 = \frac{5.25}{V} \rightarrow V = \frac{5.25}{10.5} = 0.5 \text{ mL}$

- b. The density of air at ordinary atmospheric pressure and 20°C is 1.19 g/L . What is the mass, in kilograms, of the air in a room that measures $12.5 \text{ m} \times 15.5 \text{ m} \times 8.80 \text{ m}$?

Handwritten Solution:

$d = 1.19 \text{ g/L}$

$V = 12.5 \times 15.5 \times 8.80 = 1650 \text{ L}$

$m = ? \text{ kilograms}$

Final Answer: 1968.5 kg

Handwritten Notes:

- Convert given volume, now that we have a conversion factor.
- $1 \text{ ft}^3 = 28.316 \text{ L}$
- $1550 \text{ L} \times \left(\frac{28.316 \text{ L}}{1 \text{ ft}^3} \right) = 43898.91281 \text{ L}$
- Convert grams to kg: $52230.42 \text{ g} = 52.23042 \text{ kg}$
- Final Answer: 52.23042 kg
- $20312.84 \text{ L} \times \left(\frac{1 \text{ L}}{28.316 \text{ L}} \right) = 717.3 \text{ L}$

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ML Morrison



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