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Chapter 7

Equilibria for Multiple Component Systems

- 7.2. With only a single component and phase, if only Fe₂(SO₄)₃ were present in our system, we would have C-P+2=1-1+2=2 degrees of freedom.
- 7.4. For a single-component system, there will never be negative degrees of freedom. The maximum number of phases that can exist in equilibrium is three, so the Gibbs phase rule would yield 1 3 + 2 = 0 degrees of freedom (like might occur at the triple point). For a two-component system, the maximum number of phases that can exist in equilibrium is four. There will never be negative degrees of freedom.
- 7.6. In this reaction, there are only three independent components (the amount of the fourth can be determined from the equilibrium constant) and 3 phases, two solids and a gas mixture. Therefore, there are 3 3 + 2 = 2 degrees of freedom.
- 7.8. For water: $n = \frac{pV}{RT} = \frac{(23.76 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 6.39 \times 10^{-3} \text{ mol of water is needed to ensure that there is a liquid and gas phase. This equals 0.115 grams. For methanol:$

$$n = \frac{pV}{RT} = \frac{(125.0 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 3.36 \times 10^{-2} \text{ mol of methanol is}$$
needed to ensure that there is a liquid and gas phase. This equals 1.08 grams.

7.10. Since, according to equation 7.11, $a_i = \frac{p_i}{p_i}$, and since $p_i^* = 760$ torr, we can calculate a_i :

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- 1. Understanding the eBook Physical Chemistry Ball Manual
 - The Rise of Digital Reading Physical Chemistry Ball Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Chemistry Ball Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry Ball Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Chemistry Ball Manual
 - Personalized Recommendations
 - Physical Chemistry Ball Manual User Reviews and Ratings
 - Physical Chemistry Ball Manual and Bestseller Lists
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 - o Adjustable Fonts and Text Sizes of Physical Chemistry Ball Manual
 - Highlighting and Note-Taking Physical Chemistry Ball Manual
 - Interactive Elements Physical Chemistry Ball Manual
- 8. Staying Engaged with Physical Chemistry Ball Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry Ball Manual
- 9. Balancing eBooks and Physical Books Physical Chemistry Ball Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry Ball Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry Ball Manual
 - Setting Reading Goals Physical Chemistry Ball Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry Ball Manual
 - Fact-Checking eBook Content of Physical Chemistry Ball Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
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