

A chemical reaction is a transformation of a chemical substance in which the molecular formulae are written using chemical formulae instead of words.

Chemical reactions are written using chemical formulae instead of words.

A coefficient represents the number of atoms or molecules present in a chemical reaction. A coefficient is placed directly in front of a chemical formula.

A subscript is present after the element symbol to indicate the number of atoms of that element in a molecule.

Why is it important to balance chemical equations?

Chemical reactions are balanced to ensure that the number of atoms of each element is the same on both sides of the equation. This is done by adding coefficients in front of the chemical formulae.

What does the total say is important to remember when balancing equations?

Remember, the total number of atoms of each element must be the same on both sides of the equation.

Week 8, Day 2

The number of protons in an atom cannot change (it is fixed) but a different number of neutrons can change.

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How does a positive ion form from a neutral atom? What is another name for a positive ion?

When an atom loses one or more electrons, it has more protons than it does electrons, so it is no longer neutral. This atom has a positive charge and is called a positive ion or cation.

How does a negative ion form from a neutral atom? What is another name for a negative ion?

When an atom gains one or more electrons, it has more electrons than it does protons, so it is no longer neutral. This atom has a negative charge and is called a negative ion or anion.

Atoms in Group 1A have 1 electron in their outer shell and Group 2A atoms have 2 electrons in their outer shell. To become more stable (like the noble gases) the outer shell must be full. They easily give up their valence electrons (Group 1A atoms have 1 electron in their outer shell and Group 2A atoms have 2 electrons in their outer shell).

Atoms in Group 1A have 1 electron in their outer shell and Group 2A atoms have 2 electrons in their outer shell. To become more stable (like the noble gases) the outer shell must be full. They easily give up their valence electrons (Group 1A atoms have 1 electron in their outer shell and Group 2A atoms have 2 electrons in their outer shell).

Atoms are formed by losing the first part of the element name and adding the suffix -ion. For example, sodium becomes sodium ion.

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When sodium and chlorine atoms react, the sodium atom gives up the valence electron, becoming a cation (positively charged ion), and the chlorine atom accepts the electron, becoming an anion (negatively charged ion).



When sodium and chlorine atoms react, the sodium atom gives up the valence electron, becoming a cation (positively charged ion), and the chlorine atom accepts the electron, becoming an anion (negatively charged ion).

The positively charged sodium ion and the negatively charged chloride ion are attracted to each other. The bond formed between the positively charged ion, adding the two ions together.

Physical Science Exploring Space Notetaking Guide

Answer

SJ Ball



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