

Solution Manual for Physical Chemistry 2nd Edition Ball

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4.14.

$$\left(\frac{\partial H_m}{\partial T}\right)_P = \frac{C_{P,m}}{T} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

and that we can estimate $C_{V,m}$ as follows:

- The molecule possesses five degrees of freedom, so $C_{V,m}$ is 2.5 R .
- The molecule possesses greater translational degrees of freedom, so $C_{V,m}$ is 2.5 R .
- The molecule possesses rotational freedom, so $C_{V,m}$ is 2.5 R .

4.15. From the equipartition theorem the translational energy of the N_2 is $\frac{3}{2}RT$ and $\frac{3}{2}RT$ for O_2 , so that the total $E_{\text{trans}} = 2.5RT$.

$$\text{so } \frac{1}{2} \left(\frac{\partial E_{\text{trans}}}{\partial T} \right)_P = 2.5R = 2.5 \left(\frac{8.314 \text{ J K}^{-1} \text{ mol}^{-1}}{1 \text{ kJ}} \right) = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.16. From the equipartition theorem $C_{V,m} = 2.5R = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1}$.

$$\text{Hence } \left(\frac{\partial E_{\text{trans}}}{\partial T} \right)_P = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1} = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

$$\text{Hence } \left(\frac{\partial E_{\text{trans}}}{\partial T} \right)_P = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1} = 2.08 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.17. From the equipartition theorem the translational energy of the molecule is $\frac{3}{2}RT$.

$$\text{so } \frac{1}{2} \left(\frac{\partial E_{\text{trans}}}{\partial T} \right)_P = 1.5R = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

Physical Chemistry Ball Solution Manual

RM Cervero



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