(d) The student found that 22.50 cm³ of 0.0800 moldm⁻³ Na₂B₄O₇ reacted with 25.00 cm³ of dilute hydrochloric acid.

$$Na_2B_4O_7 + 2HCl + 5H_2O \rightarrow 2NaCl + 4H_3BO_3$$

(i) Calculate the amount, in mol, of Na₂B₄O₇ used.

(ii) Calculate the amount, in mol, of HCl used.

(iii) Calculate the concentration, in moldm⁻³, of the HCL

[Total: 12]

END OF QUESTION PAPER



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- 1. Understanding the eBook Ocr 17 Jan 2013 Canswer Paper
 - o The Rise of Digital Reading Ocr 17 Jan 2013 Canswer Paper
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ocr 17 Jan 2013 Canswer Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - ∘ Features to Look for in an Ocr 17 Jan 2013 Canswer Paper
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ocr 17 Jan 2013 Canswer Paper
 - Personalized Recommendations
 - Ocr 17 Jan 2013 Canswer Paper User Reviews and Ratings
 - o Ocr 17 Jan 2013 Canswer Paper and Bestseller Lists
- 5. Accessing Ocr 17 Jan 2013 Canswer Paper Free and Paid eBooks
 - o Ocr 17 Jan 2013 Canswer Paper Public Domain eBooks

- Ocr 17 Jan 2013 Canswer Paper eBook Subscription Services
- o Ocr 17 Jan 2013 Canswer Paper Budget-Friendly Options
- 6. Navigating Ocr 17 Jan 2013 Canswer Paper eBook Formats
 - o ePub, PDF, MOBI, and More
 - o Ocr 17 Jan 2013 Canswer Paper Compatibility with Devices
 - Ocr 17 Jan 2013 Canswer Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr 17 Jan 2013 Canswer Paper
 - Highlighting and Note-Taking Ocr 17 Jan 2013 Canswer Paper
 - Interactive Elements Ocr 17 Jan 2013 Canswer Paper
- 8. Staying Engaged with Ocr 17 Jan 2013 Canswer Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr 17 Jan 2013 Canswer Paper
- 9. Balancing eBooks and Physical Books Ocr 17 Jan 2013 Canswer Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr 17 Jan 2013 Canswer Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr 17 Jan 2013 Canswer Paper
 - Setting Reading Goals Ocr 17 Jan 2013 Canswer Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr 17 Jan 2013 Canswer Paper
 - Fact-Checking eBook Content of Ocr 17 Jan 2013 Canswer Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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