

A chemical bond - forms when 2 or more atoms rearrange \_\_\_\_\_  
\_\_\_\_\_ to increase \_\_\_\_\_.

ionic bond - forms when valence \_\_\_\_\_ and \_\_\_\_\_  
from one atom to another.

cation - atom \_\_\_\_\_ electrons to become \_\_\_\_\_ charged.

anion - atom \_\_\_\_\_ electrons to become \_\_\_\_\_ charged.

In ionic compounds the ions are arranged in a \_\_\_\_\_  
\_\_\_\_\_ forces hold the ions together.

Properties of ionic compounds:

- High \_\_\_\_\_ and \_\_\_\_\_ points
- \_\_\_\_\_ - not easily \_\_\_\_\_
- \_\_\_\_\_ electricity when \_\_\_\_\_ or  
\_\_\_\_\_ because the ions are free to \_\_\_\_\_.

covalent bond - \_\_\_\_\_ and \_\_\_\_\_, forming  
\_\_\_\_\_.

Covalent compounds have \_\_\_\_\_ forces holding the  
\_\_\_\_\_ together.

Properties of covalent compounds:

- lower \_\_\_\_\_ and \_\_\_\_\_ points
- Many covalent compounds are \_\_\_\_\_ liquids or gases.
- \_\_\_\_\_ - easier to \_\_\_\_\_
- are not \_\_\_\_\_ of electricity

electronegativity - property that tells how strong an atom's \_\_\_\_\_  
is for \_\_\_\_\_.

Since oxygen has a \_\_\_\_\_ electronegativity than hydrogen, oxygen  
holds onto shared electrons \_\_\_\_\_, giving the oxygen a \_\_\_\_\_  
negative charge and the hydrogen a partial \_\_\_\_\_ charge.

# Note Taking Guide Key Episode 501

**JR Anderson**



## Note Taking Guide Key Episode 501:

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